

# OPTIMIZATION OF LEACHING RECOVERY OF ZINC, IRON AND MANGANESE FROM SELECTED SOLID WASTE ASHES IN ZARIA, KADUNA STATE, NIGERIA

## Abstract

*The increasing environmental burden of toxic heavy metal-concentrated ash from solid waste incineration, particularly from burnt tires and contraband drugs, poses significant challenges for waste management and resource recovery. This study aims to optimize the leaching recovery of zinc, iron, and manganese from these solid waste ashes in Zaria, Kaduna State, Nigeria, using Response Surface Methodology (RSM). The ashes were analyzed for their metal content using atomic absorption spectroscopy (AAS), with additional characterization by X-ray diffraction (XRD) and X-ray fluorescence (XRF). Among the leaching agents tested, hydrochloric acid (HCl) exhibited the highest efficiency, particularly under optimized conditions determined through central composite design (CCD). The optimal conditions for leaching from burnt tire ash were found to be 2.97 mol/l acid concentration, 97.7 minutes leaching time, and 68.85°C temperature, yielding recoveries of 32,009 mg/kg zinc, 4,553 mg/kg iron, and 137 mg/kg manganese. For contraband drug ash, the optimal conditions were 2.83 mol/l acid concentration, 97.7 minutes leaching time, and 68.85°C temperature, resulting in recoveries of 70.55 mg/kg zinc, 5,771 mg/kg iron, and 676 mg/kg manganese. This study successfully developed a multiple linear regression model to express the relationship between the leaching parameters and metal recovery, providing a reliable and sustainable method for metal recovery from solid waste ashes. The findings offer valuable insights into waste management practices, contributing to the circular economy and reducing the environmental impact of waste disposal.*

**Key Words:** Leaching Recovery, Zinc, Solid Waste Ashes, Response Surface Methodology (RSM)

**Hydrochloric Acid (HCl)**

## **1.0 Introduction**

The steady increase in garbage around the world is one of the most pressing concerns confronting modern society, as it is linked to global pollution (Singh and Singh 2017; Ebadi *et al.*, 2020). Environmental pollution has been regarded as one of the major global challenges in recent times (Silpa *et al.*, 2018). The idea of converting one of humanity's biggest challenges (pollutant) into a useful resource to address yet another challenge, the energy crisis and the shortage of economic heavy metals, is mind blowing (Funari *et al.*, 2015; Imran *et al.*, 2020).

Ultra-rapid industrialization, urban migration and rapid population growth have been accused of being the factors responsible for this scenario of global polluted/potentially polluted environments. This is more or less true; however, the major cause of an unhealthy environment today is simply the attitude towards it. In many developing nations, the failure to address the issue of proper waste disposal, which is the major cause of environmental pollution, has been traced to factors such as unhealthy socio-cultural practices, poor environmental sanitation habits, lack of awareness, low literacy level, bad governance, disregard to the rule of law and other forms of indiscipline (Sylvester&Ikudayisi, 2021). In addition, illicit drugs seized by the National Drug Law Enforcement Agency (NDLEA) have reportedly been burned alongside other solid waste materials, as shown in the Plate1. In the statement released, "Over 560,068.31414 kilogramme of assorted illicit drugs were on Thursday 4th August 2022 set ablaze by the National Drug Law Enforcement Agency, NDLEA, the largest to be destroyed in the 32-years history of the anti-narcotic agency" (NDLEA, 2022).

As a consequence of poorly managed waste, it pollutes water bodies, blocks drainages, causes flooding, transmits diseases through breeding of vectors, contaminates agricultural farmlands and crops, and pollutes the air through airborne particles from burning of the waste, among several other issues (Silpa *et al.*, 2018).

### **1.1 Response Surface Methodology**

In 1951, as cited by (Djimtoingar *et al.*, 2022) Box and Wilson pioneered Response Surface Methodology (RSM) as a strategy to improve chemical manufacturing processes, particularly within the chemical industry. Their primary focus was on optimizing chemical reactions, aiming to achieve favourable results such as increased yield and purity while simultaneously reducing costs. They pursued a structured approach, employing a series of sequential experiments that

took into account various variables, including temperature, pressure, reaction time, and the proportions of reactants, as described by Dean *et al.* (2017). Response Surface Methodology (RSM) stands as a robust optimization approach extensively applied for evaluating the optimal process conditions when limited experimental data is available, as highlighted by Bayuo *et al.* in 2020. This method has garnered popularity for several compelling reasons, including its ability to obviate the need for excessive chemical usage, reduce time requirements, minimize costs, and alleviate labour-intensive efforts, as emphasized by Şahan and Öztürki (2014).

Further, RSM effectively addresses the limitations of single-factor optimization, a concern articulated by Aghaeinejad-Meybodiet *al.* (2015) and Kakelar and Ebrahimi (2016). RSM encompasses a suite of mathematical and statistical techniques centered on fitting empirical models to experimental designs. To achieve this goal, linear or quadratic polynomial functions are employed to characterize the system under investigation, allowing for the exploration and adjustment of experimental conditions until they are optimized, as elaborated by Vezaet *al.* (2023). The methodology proves valuable in enhancing, designing, advancing, and refining processes in scenarios where multiple variables influence one or more responses (Ghorbannezhadet *al.*, 2016; Boateng , 2013). The term "optimization" is widely employed in analytical chemistry to denote the process of identifying the conditions that yield the most favourable response. When experimental data cannot be adequately described by linear functions, techniques for designing quadratic response surfaces, such as three-level factorial, Box-Behnken, central composite, and Doehlert designs as outlined in Sudha *et al.* (2013), are recommended for approximating a response function.

To employ RSM effectively, specific prerequisites must be satisfied. These criteria encompass:

1. Identification of the pivotal process factors: RSM relies on a prior understanding of the critical factors influencing the process. It is important to note that RSM software is typically tailored to handle a restricted number of variables, and as the number of variables grows, the volume of experiments required by the designs escalates dramatically. Consequently, most RSM programmes are generally constrained to managing four or five variables.

2. Continual variation of factors: The factors under investigation must exhibit continuous variations within the range examined, allowing for a comprehensive assessment of their impact.
3. Existence of a mathematical relationship: A mathematical function should be available that establishes a connection between the response and the factors involved, as stated by Geiger in 2014.

Metals are rarely found in the elemental forms in their natural occurring state in the earth's crust, except a few such as gold. Many metals, however, are found in mineral forms, in combination with one or many others in a form called ores. Zinc element is usually found in combination with lead or copper. Sphalerite (ZnS), also known as zinc blende, is regarded as the most abundant zinc mineral (Srivastava and Tanveer, 2020).

### **1.3 Iron ores and metal**

Iron is the second most abundant element in the Earth's crust, after aluminium; and fourth in abundance behind oxygen, silicon, and aluminium among the rest of the elements, making up about 5% of the Earth's crust (Geerdes *et al.*, 2020). However, it is the most abundant element in the earth as a whole, accounting for about 35%, and is relatively plentiful in the Sun and other stars.

Iron rarely exists in pure form in the crust. However, it exists as terrestrial iron (alloyed with 2–3 percent nickel) in basaltic rocks in Greenland and carbonaceous sediments in the United States (Missouri) and as low-nickel meteoric iron (5–7 percent nickel), kamacite (Encyclopaedia Britannica, 2022). Nickel-iron, a native alloy, occurs in terrestrial deposits (21–64 percent iron, 77–34 percent nickel) and in meteorites as taenite (62–75 percent iron, 37–24 percent nickel). Also, iron is found combined with other elements in hundreds of minerals of greatest importance are as iron ore, some of which include hematite (ferric oxide,  $\text{Fe}_2\text{O}_3$ ), magnetite (triiron tetroxide,  $\text{Fe}_3\text{O}_4$ ), limonite (hydrated ferric oxide hydroxide,  $\text{FeO}(\text{OH})\cdot n\text{H}_2\text{O}$ ), and siderite (ferrous carbonate,  $\text{FeCO}_3$ ) (Encyclopaedia Britannica, 2022).

#### **1.4 Manganese ore and metal**

Manganese (Mn) is a naturally occurring metallic element and is the twelfth most abundant metallic element in the Earth's crust; constituting 0.1% of the earth's crust, depending on the model used for the crustal composition (Barry, 2010; Baba *et al.*, 2014). It is usually found in rock, soil and water. Manganese rarely exists in a pure form; instead, it combines with other elements in nearly 300 different minerals. It is extracted from several economically important oxide minerals, including pyrolusite ( $\text{MnO}_2$ ), psilomelane ( $\text{BaMn}_9\text{O}_{16}(\text{OH})_4$ ), manganite ( $\text{Mn}_2\text{O}_3\text{H}_2\text{O}$ ), and wad (Baba *et al.*, 2014).

Over 80% of the world's manganese deposits are found in South Africa and Ukraine, while other important manganese deposits are in China, Australia, Brazil, Gabon India, and Mexico. The economically mineable manganese deposits occur in the natural concentrations of 150-500 times the average crustal abundance (Baba *et al.*, 2014). Geological processes have created natural concentrations of mineable grade and sizes in various places, formed by distinct geological processes due to the mobility of the element (Baba *et al.*, 2014). The decrease in the reserves of high-grade manganese ores is making it necessary to consider the processing of lower-grade manganese ores and sources (Spooren et al., 2020).

#### **1.5 Municipal Solid Waste Incineration (MSWI) Ash**

Municipal solid waste (MSW) incineration ash is basically product of the combustion of municipal solid waste combustion or its incineration, which is a high-temperature dry oxidation process, that converts flammable and organic waste to non-flammable and inorganic waste (Ako& Ma, 2022). MSW incineration ash is mainly constituted by two main forms of ash: bottom ashes (BsA) and fly ashes (FA) (Tang, 2017). As the words suggest, bottom ash is collected at the bottom of the combustion chambers and is the major constituent of combustion ash. On the other hand, fly ash is collected in the flue gas cleaning system of the combustion plant. The former contains materials like metal objects, slag, glassy silicate and minerals with high melting points.

Meanwhile, fly ash contains small dust particles of varying sizes (needle, strip and spherical), high porosity, low water content, large specific surface area, much lighter weight, and relatively low melting points (Ako& Ma, 2022). In any case, both the bottom ash and fly ash, but

particularly fly ash contain varied but considerable amounts of metals, salts, organic pollutants, among other components, some of which might be very toxic (Bakaláet *al.*, 2020). Some typical fly ashes are characterized by high chloride content and significant amount of heavy metals and toxic organic compounds, especially polychlorinated dibenzo-p-dioxins and polychlorinated dibenzofurans (PCDDs and PCDFs) (Ako& Ma, 2022). Generally speaking, fly ash is composed mainly of soluble salts such as CaO, SiO<sub>2</sub>, Al<sub>2</sub>O<sub>3</sub>, Fe<sub>2</sub>O<sub>3</sub>, MgO, Na<sub>2</sub>O, K<sub>2</sub>O, P<sub>2</sub>O<sub>5</sub>, TiO<sub>2</sub>, MnO, CuO, ZnO, PbO, Cr<sub>2</sub>O<sub>3</sub>, SO<sub>3</sub>, and by a minor fraction of other components such as Zn, Pb, Cr, Cd, Cu, Sn, Ba, Sb, Zr, As, Co, Mo, Rb, Bi, V, Ce, Ga, La, Nd, Nb, and Hg (Ako& Ma, 2022). Therefore, the leaching out of these heavy metals and their impact on human health and environmental pollution have become major concerns, making it imperative to carefully investigate and select a safe and efficient method of treating fly ash.

## **2.0 MATERIALS AND METHODOLOGY**

### **2.1 Materials**

#### **2.1.1 Apparatus, Equipment, and Reagents.**

The following are the apparatus and equipment used for the research work: Watch glass, funnel, retort stand, thermometer, oven, magnetic stirring machine, electronic balance (Sartorius Extend ED224S, Germany), ball milling machine (PMT/500/250/HD, India), pulverizing machine (Yam 1100 Q 4, China), sieve shaker, x-ray fluorescence (XRF) machine (X-MET8000 spectrophotometer, England), x-ray diffractometer (XRD) (RikaguMiniflex, Japan), atomic absorption spectrophotometer (AAS) machine (AgilentTechnologies200 Series AA (280 FS AA), USA). Furthermore, all reagents used for this research were of analytical grades. Additionally, distilled water was used for the preparation of all solutions. The leaching reagents used for the experiments include hydrochloric acid, tetraoxosulphate (VI) acid, trioxonitrate (V) acid, citric acid, oxalic acid, ammonia, sodium hydroxide and deionized water. All the glassware was cleaned with distilled water and dried in an oven before use.

### **2.2 Methodology**

### **2.2.1 Collection and Preparation of Ash Samples**

The ash samples marked BT (burnt tyre ash) and AS (contraband drugs ash) were sourced based on activity on-site from two dump sites located at Dogarawa, Sabon Gari Local Government, Zaria, Kaduna, Nigeria; after an initial survey on several municipal solid waste dump sites at different locations in Zaria. BT and AS were samples of burnt tyre ash and contraband drug ash respectively. Each was collected as composite samples, obtained from five points (diagonals and the centre of each of the dump sites; 60 cm apart) at weekly intervals for four weeks and merged into bulk samples in November, 2021. From each collecting point, about 500 g of ash was collected using a clean glass beaker.

The bulk samples were prepared for analysis and experimentation with slight modifications of the method described by Baba and Adekola (2011). The sample of each ash was oven-dried at 105°C for 4 hours and ground into a fine powder using a pulverizing device. The ash of each group was then sieved into three fractions of particle size (75 µm, 150 µm, and 250 µm) using the ASTM standard sieve at the Department of Geology, Ahmadu Bello University, Zaria. Unless otherwise stated, characterization of the ashes and leaching experiments were carried out on the smallest fraction.

### **2.2.2 The Leaching of Metals from Ashes**

Metals of economic value (Zn, Fe, and Mn) were selectively leached from each of the BT and AS samples. The procedure was carried out by using several acids and alkaline leaching agents as suggested by Kinoshita *et al.*, 2005. Hydrochloric acid (HCL), trioxonitrate(V) acid (HNO<sub>3</sub>), tetraoxosulphate(VI) acid (H<sub>2</sub>SO<sub>4</sub>), citric acid, oxalic acid, sodium hydroxide (NaOH), ammonia (NH<sub>3</sub>), and deionized water were used as leaching agents. The BT and AS samples (1.0 g each) were separately dissolved in 100 cm<sup>3</sup> of each leaching agent (1.0 M) in separate 250 cm<sup>3</sup> glass beaker with watch glass lid. The leaching process was conducted at a temperature of 80°C for 2 hours at a stirring speed of 360 rpm. Afterward, the metal concentrations in the leached liquors of each groups were measured by AAS after filtering out the solid residues. The leaching agent with the highest leaching efficiency and the corresponding metals of economic value was considered for use in the optimization study.

### **2.2.3 Characterization of Sample**

The elemental composition, as well as the mineralogical phase identification of the raw samples of BT and AS was determined using X-ray fluorescence spectrometry and X-ray diffractometry. The aqueous metal analysis of the BT and AS samples was carried out using Atomic Absorption Spectrophotometry (AAS) to determine the concentration of each of the metals present in the selected aqueous or leached liquor, during the leaching and optimization of the leaching experiments.

## 2.3 Analysis

### 2.3.1 Multiple linear regression model for the recovery of metals from BT and AS samples

Design Expert statistical software Version 12 (STAT-EASE Inc., Minneapolis, USA) was used for statistical calculations, modelling, and plotting. Based on the literature and initial experiments, the effects of three parameters, including concentration (0.5-12.0 Mol/dm<sup>3</sup>), temperature (25-80°C), and time (10-120 min), on the leaching of the three selected metals of economic interest from BT and AS ashes collected from Dogarawa area in Zaria, Nigeria were considered.

The RSM method is suitable for fitting a quadratic surface and aids in the analysis of parameter interactions, as well as the optimization of process parameters with the least number of experiments (Behera *et al.*, 2018). Three responses (Zn, Fe, and Mn) were chosen to be optimized in a nonlinear optimization, where the experimental data were fitted to a second polynomial equation as shown in equation (3.2), and regression coefficients were obtained. However, the software randomizes the experimental run to diminish the error and the effect of uncontrolled factors (Anderson and Whitcomb, 2016).

$$\gamma = \beta_0 + \sum_{i=1}^n \beta_i X_i + \sum_{i=1}^n \beta_{ii} X_i^2 + \sum_{i=1}^n \sum_{j=i+1}^n \beta_{ij} X_i X_j + \varepsilon \quad (3.2)$$

Where  $\gamma$  is the predicted response;  $\beta_0$  the constant coefficient;  $\beta_i$  the linear coefficients;  $\beta_{ii}$  the quadratic coefficients;  $\beta_{ij}$  the interaction coefficients;  $n$  the number of factors studied and optimized in the experiments;  $X_i$  and  $X_j$  are the coded values of the variable parameters for leaching process; and  $\varepsilon$  is the random error.

### 2.3.2 Statistical analysis

The analysis of variance (ANOVA) was performed to justify the significance and adequacy of the developed regression model. The adequacy of the response surface model was evaluated by calculating the determination coefficient ( $R^2$ ) and also by testing it for the lack of fit. The significance of model terms was evaluated by the probability value (P-value) at a 95% confidence interval.

## 3.0 Results and Discussion of Findings

### 3.1 Quality Assurance

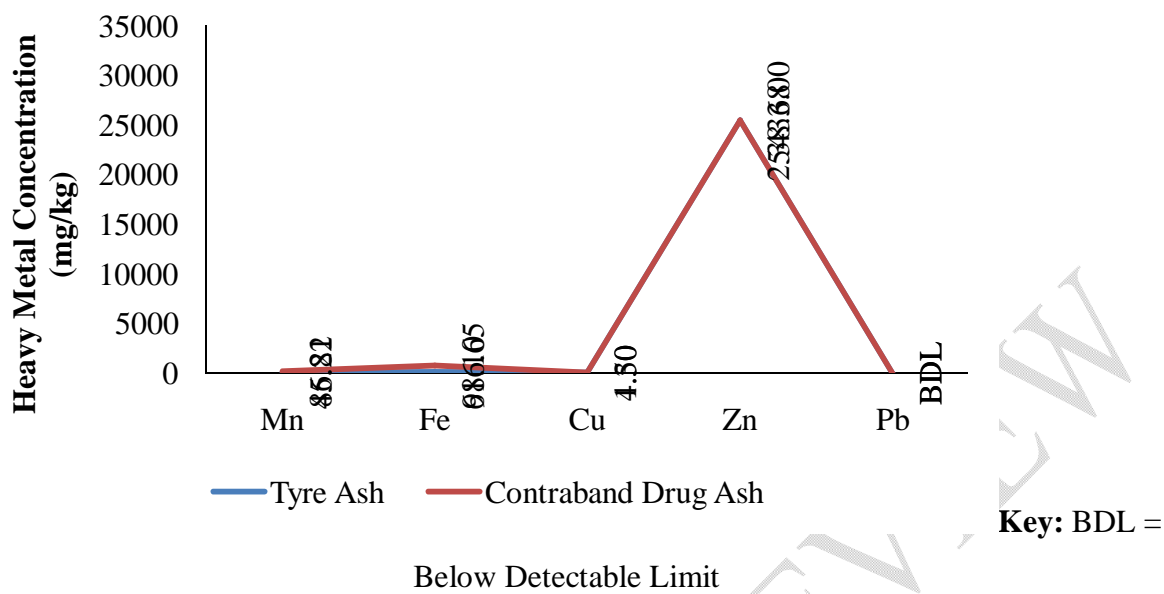
Table 1 Reveals the percentage recovery of Zn, Fe and Mn from BT and AS ash samples analysed from the spiking experiment.

**Table.1: Percent recovery of Zn, Fe and Mn from BT and AS ash samples**

%Recovery		
Metal	BT Sample	AS Sample
Zn	95.34	86.34
Fe	89.60	90.15
Mn	81.04	82.90

#### 3.2.1 Determination of heavy metals present in ash samples

The first analysis carried out on the prepared ash samples aimed to determine the type(s) and amount of heavy metals present in the two ash samples. The five (5) heavy metals tested for, in both the BT and AS sample is presented in Figure 1.



**Figure 1: Heavy metal concentration in BT and AS**

Figure 1 shows that the concentration of Mn in burnt tyre ash was  $45.22 \pm 0.58$  mg/kg, while those of Fe, Cu and Zn were  $98.10 \pm 2.21$ ,  $1.30 \pm 0.02$  and  $25433.00 \pm 733.76$  mg/kg respectively. Analyses of the contraband drug ash revealed that concentration of Mn was  $86.81 \pm 2.56$  mg/kg, while those of Fe, Cu, and Zn were  $616.65 \pm 53.82$ ,  $4.50 \pm 0.11$  and  $38.68 \pm 0.63$  mg/kg respectively. The concentration of Pb, however, was below the detectable limit in both ash samples as presented in table 2 and 4.2.2

**Table 2: Preliminary determination of heavy metals in burnt tyre ash**

S/No.	Heavy metal	1 <sup>st</sup>	2 <sup>nd</sup>	3 <sup>rd</sup>	Average	SD
1.	Mn (mg/kg)	45.43	44.42	45.80	45.22	0.58
2.	Fe (mg/kg)	95.23	100.61	98.46	98.10	2.21
3.	Cu (mg/kg)	1.30	1.27	1.33	1.30	0.02
4.	Zn (mg/kg)	24400.00	26035.00	25864.00	25433.00	733.77
5.	Pb(mg/kg)	BDL	BDL	BDL	N/A	N/A

Key: BDL: Below Detectable Limit, N/A: Not Applicable

**Table 3: Preliminary determination of heavy metals in contraband drug ash**

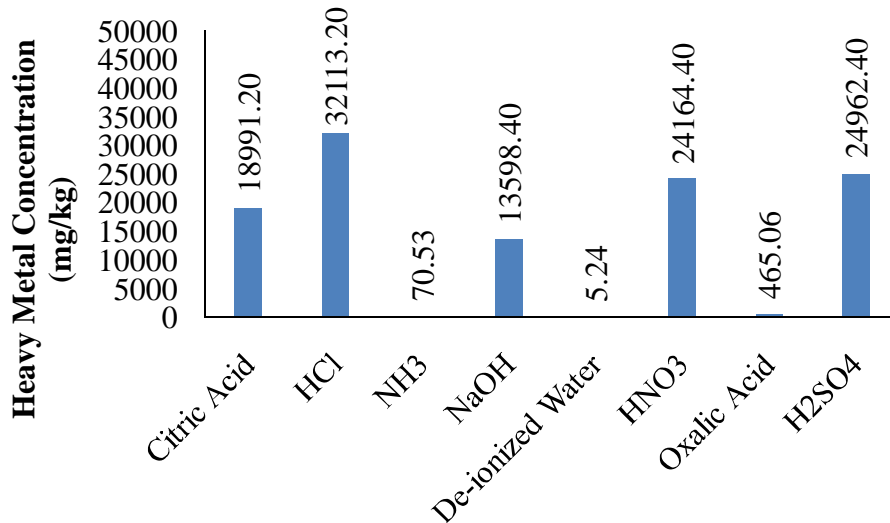
S/No.	Heavy metal	1 <sup>st</sup>	2 <sup>nd</sup>	3 <sup>rd</sup>	Average	SD
1.	Mn (mg/kg)	83.62	89.91	86.89	86.81	2.57
2.	Fe (mg/kg)	688.62	559.2	602.13	616.65	53.82
3.	Cu (mg/kg)	4.62	4.53	4.35	4.50	0.11
4.	Zn (mg/kg)	39.43	37.89	38.72	38.68	0.63
5.	Pb (mg/kg)	BDL	BDL	BDL	N/A	N/A

**Key:** BBL: Below Detectable Limit, N/A: Not Applicable

From the results, it was observed that the concentration of Zn, Fe, and Mn were the highest in the two ash samples. Consequently, these heavy metals were selected for further investigation.

### 3.2.2 Selection of leaching reagent

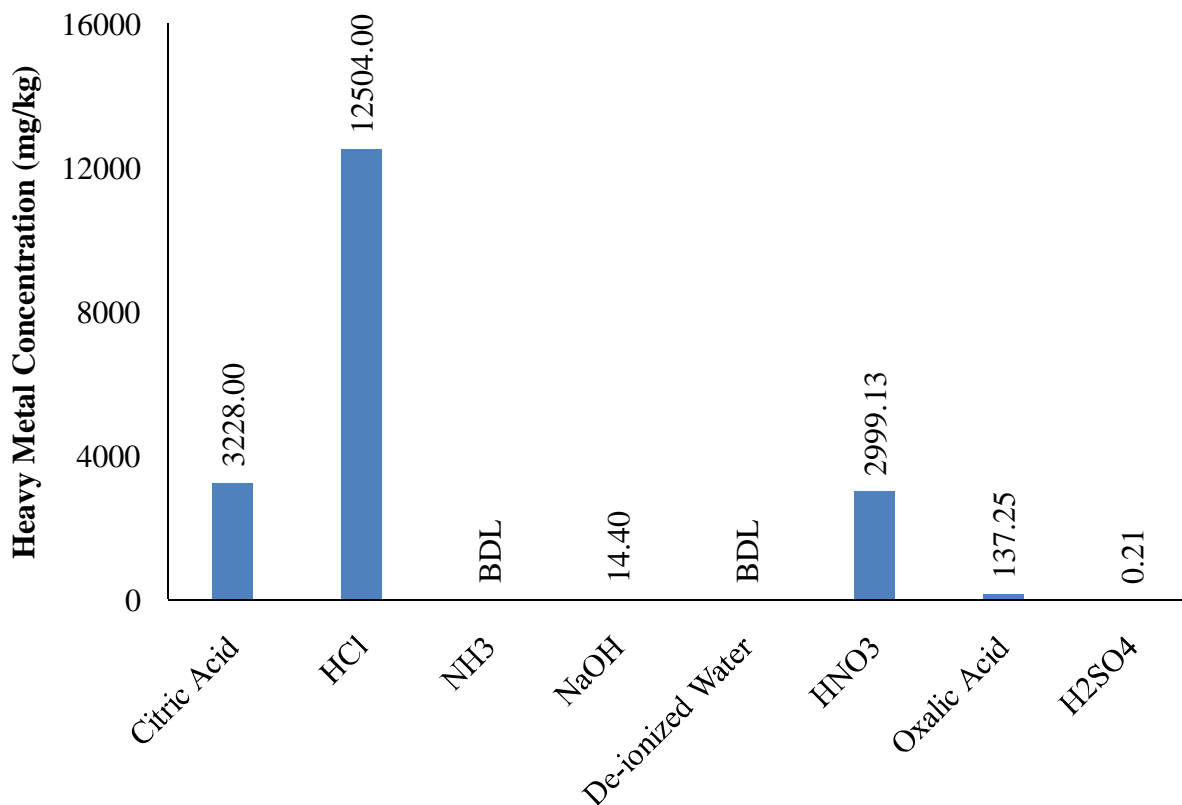
Three (3) heavy metals (Zn, Fe, and Mn) were selected for further investigation. In this section, the leaching behaviour of these metals from the two selected solid waste samples, BT and AS, was investigated using eight (8) leaching agents. The result of the use of the eight reagents to leach the three heavy metals from the ash samples is presented in Figure.2.



**Figure 2: Concentration of Zn leached from burnt tyre ash by various reagents**

From the findings, the concentration of Zn leached from the burnt tyre ash by each of the eight reagents indicated that HCl, H<sub>2</sub>SO<sub>4</sub>, and HNO<sub>3</sub> leached 32,113.20 ± 892.83, 24,962.40 ± 1507.44, and 24,164.40 ± 1308.54 mg/kg of Zn, respectively. Meanwhile, citric acid and NaOH leached 18,991.20 ± 892.83 and 13,598.40 ± 1380.65 mg/kg of Zn, respectively.

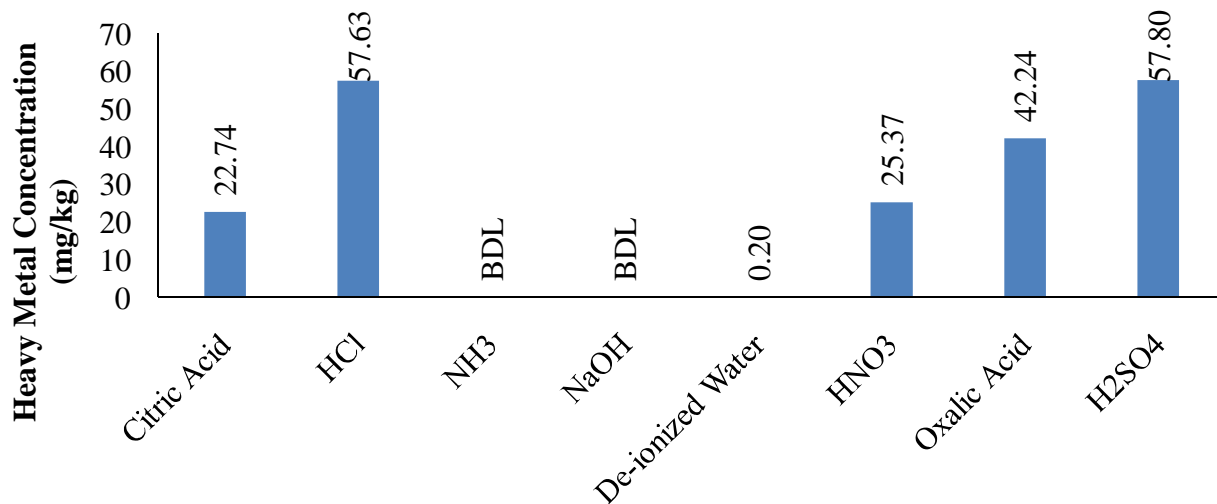
The concentration of Fe leached from the BT by the reagents is presented in Figure.3.



**Figure 3: Concentration of Fe leached from burnt tyre ash by various reagents**

The results revealed that HCl, citric acid and HNO<sub>3</sub> leached 12,504.00 ± 1393.03, 3,228.00 ± 206.56 and 2,999.13 ± 118.73 mg/kg of Fe respectively; while oxalic acid and NaOH leached 137.25 ± 6.14 and 14.4 ± 1.04 mg/kg of Fe respectively.

The concentration of Mn leached from BT by the reagents is presented in Figure 4.



**Figure 4: Concentration of Mn leached from burnt tyre ash by various reagents**

The results revealed that H<sub>2</sub>SO<sub>4</sub>, HCl, and oxalic acid leached  $57.80 \pm 4.26$ ,  $57.63 \pm 5.18$ , and  $42.24 \pm 1.23$  mg/kg of Mn respectively; while HNO<sub>3</sub> and citric acid leached  $25.37 \pm 0.33$  and  $22.74 \pm 3.15$  mg/kg of Mn, respectively as presented in table 4. and 4.2.8

**Table 4.: Amount of Zn leached from burnt tyre ash by reagents**

S/No.	Reagent	Reading (mg/kg)			Average	SD
		1st	2nd	3rd		
1	Citric Acid	19037.5	17875.3	20060.8	18991.20	892.83
2	HCl	32547.2	34123.8	29668.6	32113.20	1844.54
3	NH <sub>3</sub>	67.62	84.65	59.311	70.53	10.55
4	NaOH	15256.6	11876.5	13662.1	13598.40	1380.65
5	De-ionized water	4.578	5.883	5.261	5.24	0.53
6	HNO <sub>3</sub>	24659.3	25461.2	22372.7	24164.40	1308.54
7	Oxalic acid	453.26	472.85	469.07	465.06	8.49
8	H <sub>2</sub> SO <sub>4</sub>	26527.5	22926.3	25433.4	24962.40	1507.44

**Table 5: Amount of Fe leached from burnt tyre ash by reagents**

S/No.	Reagent	Reading (mg/kg)			Average	SD
		1 <sup>st</sup>	2nd	3rd		
1	Citric Acid	2986.40	3491.00	3206.60	3228.00	206.56
2	HCl	12595.90	14162.30	10753.80	12504.00	1393.03
3	NH <sub>3</sub>	0	0	0	0	0.00
4	NaOH	13.30	15.80	14.10	14.40	1.04
5	De-ionized water	0	0	0	0	0.00
6	HNO <sub>3</sub>	2890.80	3164.40	2942.20	2999.13	118.73
7	Oxalic acid	139.34	143.50	128.91	137.25	6.14
8	H <sub>2</sub> SO <sub>4</sub>	0	0.42	0.21	0.21	0.17

**Table 6: Amount of Mn leached from burnt tyre ash by reagents**

S/No.	Reagent	Reading (mg/kg)			Average	SD
		1st	2nd	3rd		
1	Citric Acid	22.44	19.04	26.74	22.74	3.15
2	HCl	56.68	64.39	51.82	57.63	5.18
3	NH <sub>3</sub>	0	0	0	0	0.00
4	NaOH	0	0	0	0	0.00
5	De-ionized water	0.27	0	0.33	0.20	0.14
6	HNO <sub>3</sub>	25.71	25.47	24.92	25.37	0.33
7	Oxalic acid	40.82	43.81	42.09	42.24	1.23
8	H <sub>2</sub> SO <sub>4</sub>	53.39	63.56	56.45	57.80	4.26

**Table 7: Amount of Zn leached from contraband drug ash by reagents**

S/No.	Reagent	Reading (mg/kg)			Average	SD
		1 <sup>st</sup>	2nd	3rd		
1	Citric Acid	41.04	36.15	40.06	39.08	2.11
2	HCl	63.42	53.31	57.68	58.14	4.14
3	NH <sub>3</sub>	13.66	11.21	16.53	13.80	2.17
4	NaOH	33.12	35.02	36.38	34.84	1.34
5	De-ionized water	0.00	0.00	0.00	0.00	0.00
6	HNO <sub>3</sub>	121.63	119.21	117.72	119.52	1.61
7	Oxalic acid	49.35	46.78	47.67	47.93	1.06
8	H <sub>2</sub> SO <sub>4</sub>	40.42	36.94	38.89	38.75	1.43

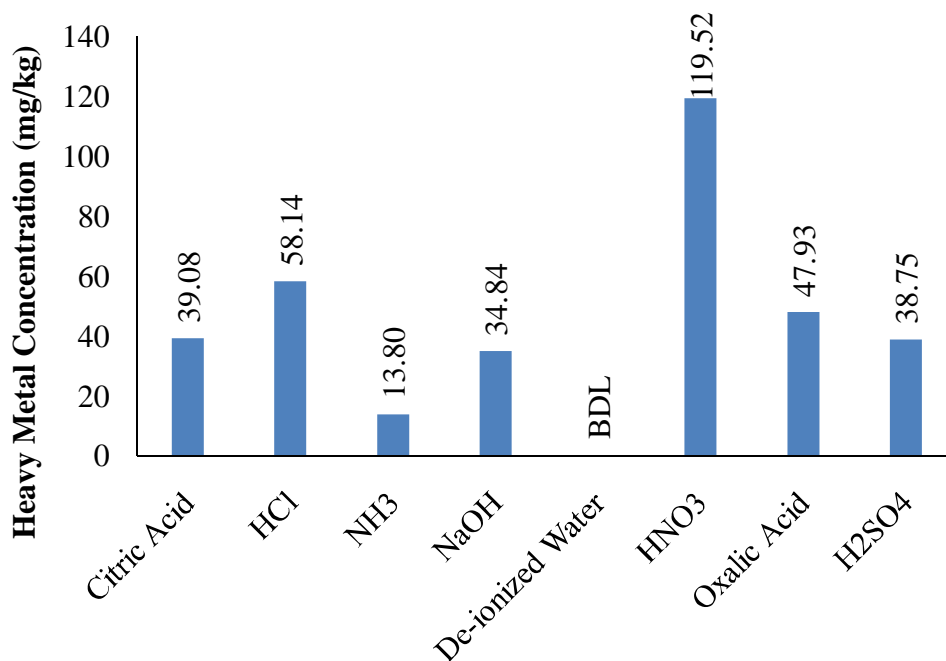
**Table 8: Amount of Fe leached from contraband drug ash by reagents**

S/No.	Reagent	Readings			Average	SD
		1 <sup>st</sup>	2 <sup>nd</sup>	3 <sup>rd</sup>		
1	Citric Acid	1435.60	1381.70	1157.10	1324.80	120.61
2	HCl	7834.60	7692.90	6982.60	7503.37	372.75
3	NH <sub>3</sub>	0.00	0.00	0.00	0.00	0.00
4	NaOH	7.86	6.39	7.12	7.12	0.60
5	De-ionized water	0.00	0.00	0.00	0.00	0.00
6	HNO <sub>3</sub>	7836.70	7759.40	8307.90	7968.00	242.41
7	Oxalic acid	6836.70	7349.20	7086.50	7090.80	209.25
8	H <sub>2</sub> SO <sub>4</sub>	0.00	0.00	0.00	0.00	0.00

**Table 9: Amount of Mn leached from contraband drug ash by reagents**

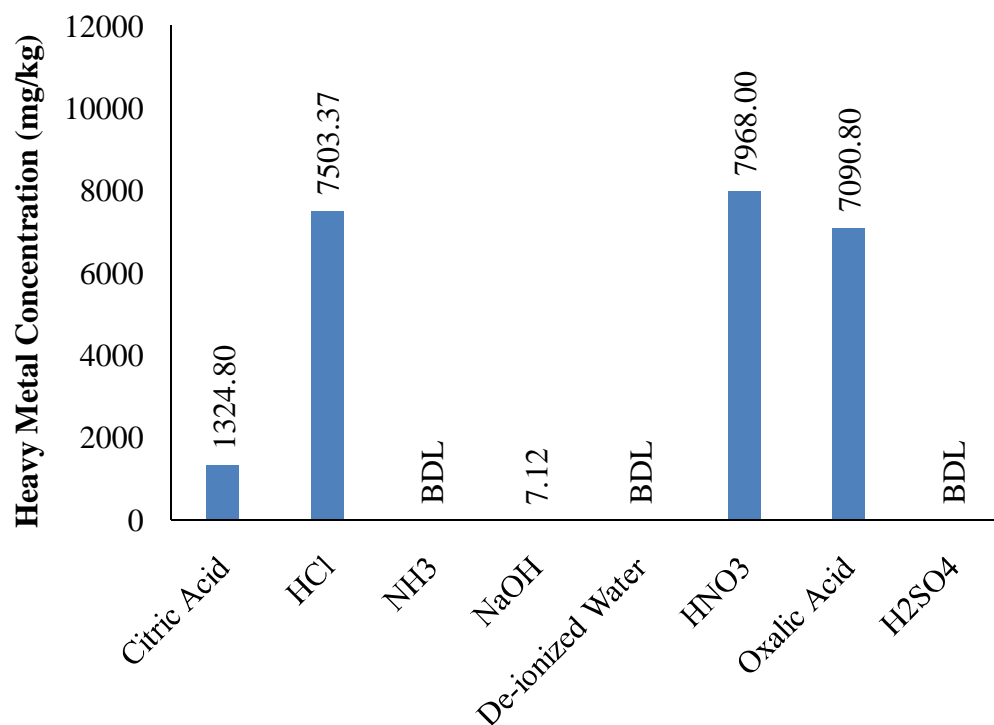
S/No.	Reagent	Readings			Average	SD
		1 <sup>st</sup>	2 <sup>nd</sup>	3 <sup>rd</sup>		
1	Citric Acid	47.67	39.86	41.06	42.86	3.43
2	HCl	74.84	80.34	72.15	75.78	3.41
3	NH <sub>3</sub>	0.00	0.00	0.00	0.00	0.00
4	NaOH	0.00	0.00	0.00	0.00	0.00
5	De-ionized water	0.00	0.00	0.00	0.00	0.00
6	HNO <sub>3</sub>	128.75	125.90	134.24	129.63	3.46
7	Oxalic acid	4.17	6.25	4.68	5.03	0.89
8	H <sub>2</sub> SO <sub>4</sub>	129.84	139.48	133.52	134.28	3.97

Figure.5 presents the amount (in mg/kg) of Zn leached from AS. The results revealed that HNO<sub>3</sub> leached 119.52 ± 1.61 mg/kg of Zn. Further, HCl, oxalic acid, citric acid, H<sub>2</sub>SO<sub>4</sub> and NaOH leached 58.14 ± 4.14, 47.93 ± 1.06, 39.08 ± 2.11, 38.75 ± 1.43, and 34.84 ± 1.34 mg/kg respectively.



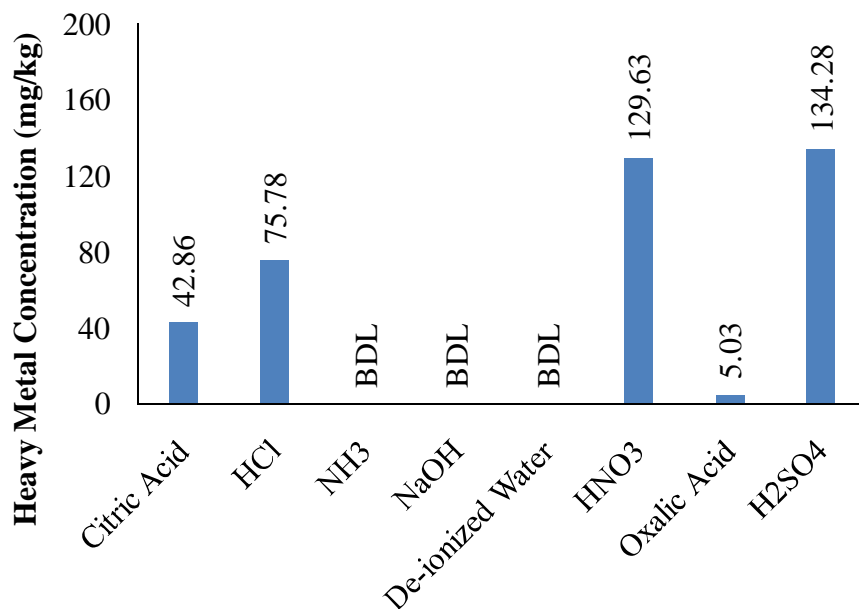
**Figure 5: Concentration of Zn leached from contraband drug ash by various reagents**

The amount of Fe leached by the reagents is as presented in Figure 6. HNO<sub>3</sub>, HCl, oxalic acid and citric acid leached 7968.00 ± 242.41, 7503.37 ± 372.75, 7090.80 ± 209.25 and 1324.80 ± 120.61 respectively.



**Figure.6: Concentration of Fe leached from contraband drug ash by various reagents**

The concentration (in mg/kg) of Mn leached from AS is shown in Figure 7.  $\text{H}_2\text{SO}_4$  and  $\text{HNO}_3$  leached  $134.28 \pm 3.97$  and  $129.63 \pm 3.46$  mg/kg of Mn respectively; while HCl and citric acid leached  $75.77 \pm 3.41$  and  $42.86 \pm 3.43$  mg/kg of Mn respectively.



**Figure.7: Concentration of Mn leached from contraband drug ash by various reagents**

### 3.3 Design of Experiments

The experimental factor levels used in the factorial design of the Response Surface Methodology (RSM) for the leaching of BT and AS, showing the range and levels of independent variables (concentration, time, and temperature) are displayed in Table 10.

**Table 10: Experimental factor levels used in factorial design**

Factor	Name	Range		Coded levels	
		Min.	Max.	-1	+1
A	Concentration (mol/l)	0.50	12.00	2.83	9.67
B	Time (min.)	10.00	120.00	32.30	97.70
C	Temperature (°C)	25.00	80.00	36.15	68.85

### 3.3.1 Coefficient determination values and linear correlation between the observed and predicted values for Zn, Fe and Mn responses for leaching of the Ash samples

Table 11 and Table 12 present the R-squared table, including Adequate Precision,  $R^2$ , Predicted  $R^2$  and Adjusted  $R^2$  for the responses of zinc, iron and manganese in the leaching of the BT and AS. Additionally, Figures 8,9 and 10 illustrate the linear correlation between the observed and predicted values for the zinc, iron and manganese responses in the leaching of BT; while Figures 11, 12 and 13 depict the same correlation for the leaching of the AS sample.

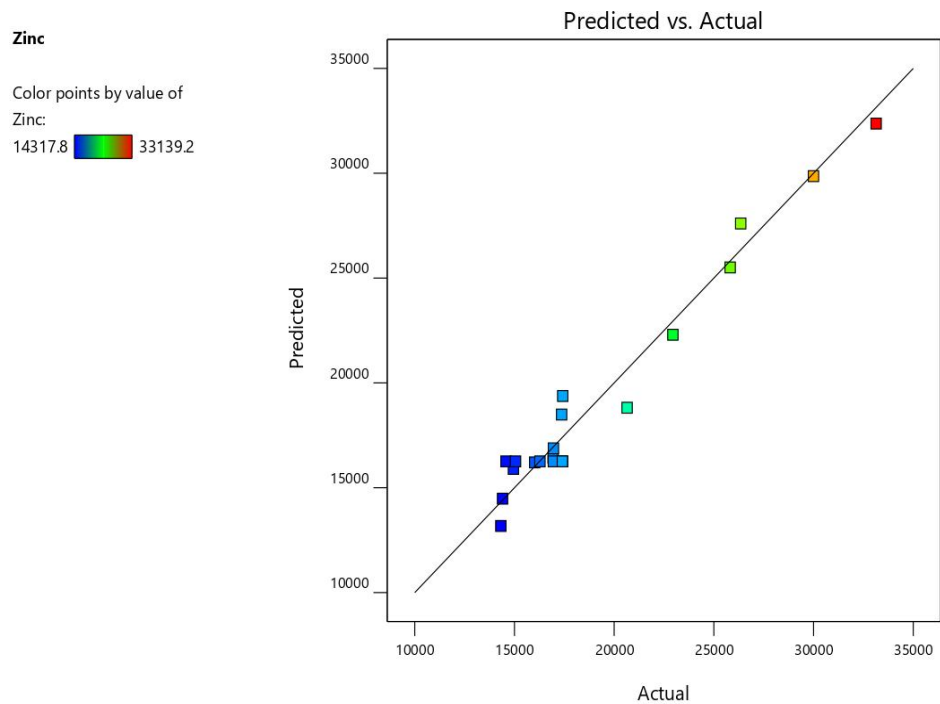
**Table 11: R-square table for zinc, iron, and manganese of the BT sample**

Fit Statistics	Zinc	Iron	Manganese
Std. Dev.	1452.33	282.43	13.67
Mean	19248.11	3258.14	96.66
C.V. %	7.54	8.6	14.14
$R^2$	0.9632	0.9494	0.9441
Adjusted $R^2$	0.9301	0.9039	0.8938
Predicted $R^2$	0.7973	0.7268	0.7021
Adeq Precision	18.7008	16.1324	18.2434

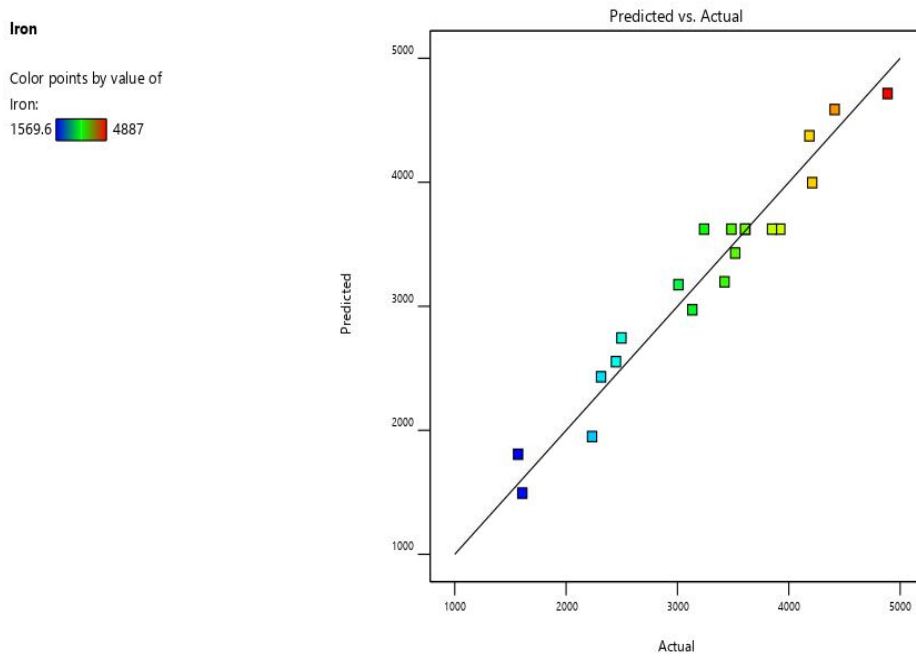
**Table 12: R-square table for zinc, iron and manganese of the AS sample**

Fit Statistics	Zinc	Iron	Manganese
Std. Dev.	1.48	227.92	40.46
Mean	47.55	3274.27	356.17

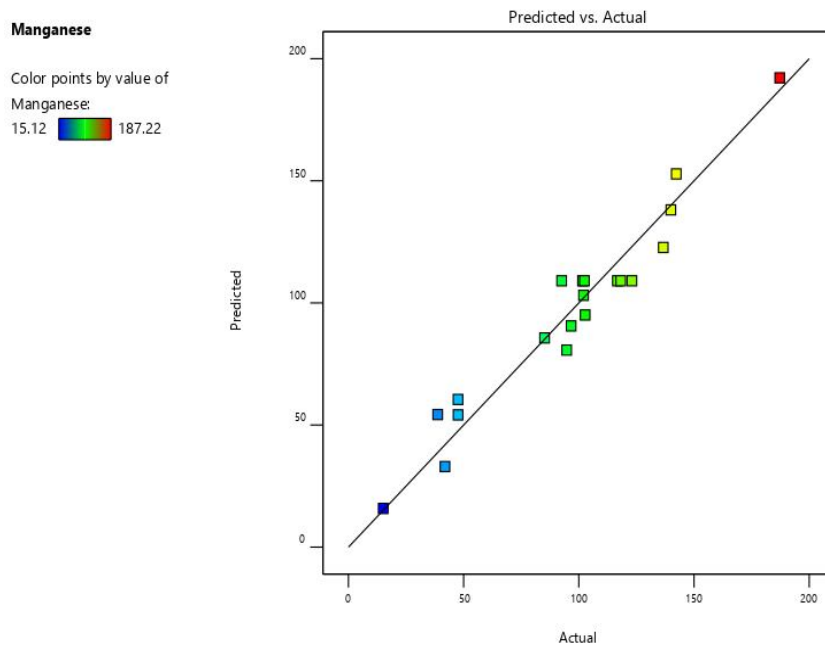
<b>C.V. %</b>	3.11	6.96	11.36
<b>R<sup>2</sup></b>	0.9893	0.9823	0.9583
<b>Adjusted R<sup>2</sup></b>	0.9797	0.9664	0.9209
<b>Predicted R<sup>2</sup></b>	0.9310	0.8902	0.7694
<b>Adeq Precision</b>	18.7008	39.7688	17.8763



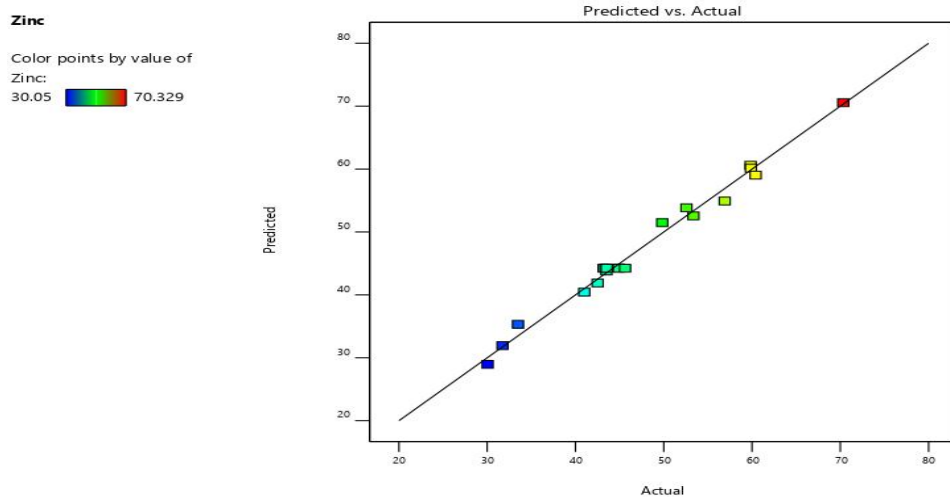
**Figure 8: Linear correlation between observed and predicted value response for leaching zinc from the BT sample**



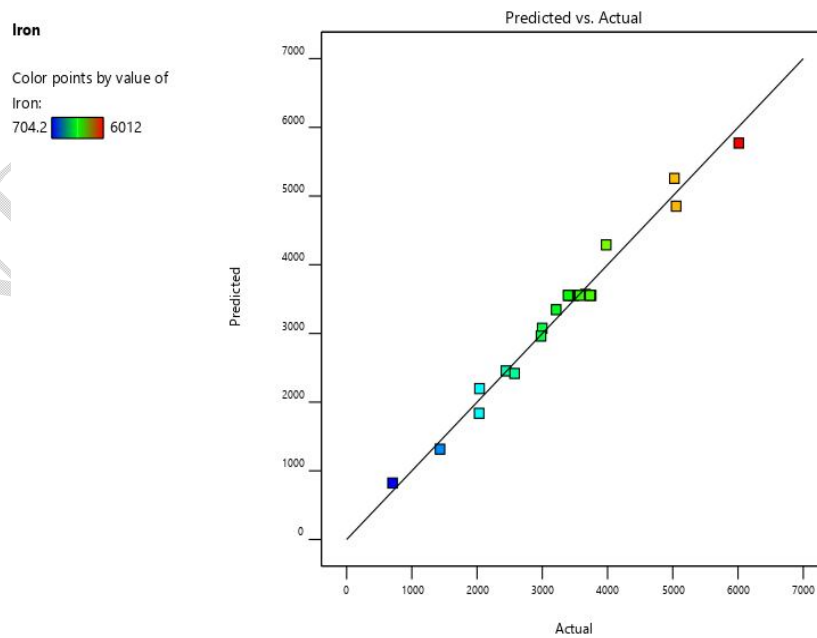
**Figure 9: Linear correlation between observed and predicted value response for leaching iron from the BT sample**



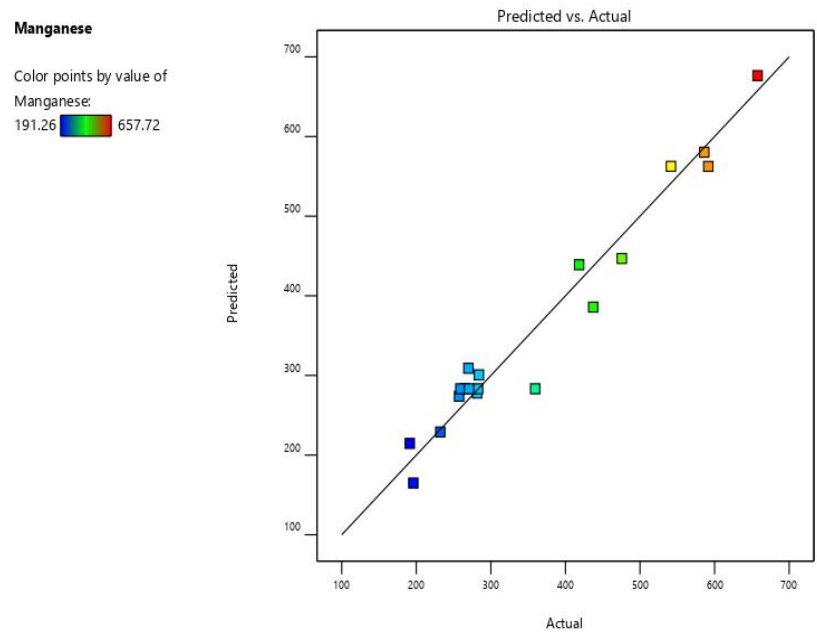
**Figure 10: Linear correlation between observed and predicted value response for leaching manganese from the BT sample**



**Figure 11: Linear correlation between observed and predicted value response for leaching of zinc from the AS sample**



**Figure 12: Linear correlation between observed and predicted value response for leaching of iron from the AS sample**



**Figure 13: Linear correlation between observed and predicted value response for leaching of manganese from the AS sample**

#### 4.0 Conclusion

The successful optimization of the leaching process using response surface methodology and the identification of optimal conditions for the recovery of zinc, iron and manganese demonstrate the potential for efficient extraction of these valuable metals from solid waste. This study provides valuable insights for the development of sustainable and economically viable strategies for the utilization of waste materials and the recovery of valuable resources. Further research and application of these findings could contribute to environmental sustainability and the development of a circular economy by reducing waste generation and promoting resource recovery.

The regression model equations developed after successfully performing the design of the experiment through CCD of the RSM indicated that the model was statistically significant, underscoring its suitability for predicting the leaching performance (responses).

## 5.0 Recommendations

Based on the results and conclusions reached in this study, the following recommendations are therefore offered for future research and application.

- i. Kinetic modeling of the process should be undertaken: Conducting kinetic modeling studies can help elucidate the leaching mechanisms and provide insights into the rate-controlling steps during the leaching process. By developing mathematical models, the leaching kinetics of zinc, iron, and manganese from solid wastes can be better understood, leading to improved process design.
- ii. Further study to investigate the effects of impurities on the leaching process should be carried out. Understanding the role of impurities will contribute to more accurate process optimization and ensure production of high-quality metal concentrates.
- iii. Further study should scale-up the work. It is essential to investigate the scalability of the optimized leaching process under large-scale conditions, simulating industrial operations. This will help determine the feasibility and potential challenges associated with implementing the optimized process on a commercial scale.
- iv. A techno-economic analysis of the large-scale process should be carried out for evaluating the economic viability of the optimized leaching process. This analysis will provide a realistic assessment of the economic feasibility and potential profitability of implementing the optimized process in a commercial setting.
- v. An environmental impact assessment of the large-scale leaching process should be carried out for the potential environmental implications due to the use of chemicals and generation of waste streams. This should include assessing the potential risk associated with the use and disposal of leaching agents, evaluating the waste management strategies and identifying potential mitigation measures to minimize the environmental footprint of the process.

By undertaking these further studies, researchers can enhance the understanding of the leaching process for zinc, iron and manganese recovery from solid waste, improve process efficiency and address potential challenges. The outcome of these studies will contribute to the development of sustainable and economically viable strategies for metal recovery from solid waste in Zaria, Kaduna State, Nigeria and beyond.

#### DISCLAIMER (ARTIFICIAL INTELLIGENCE)

Author(s) hereby declares that NO generative AI technologies such as Large Language Models (ChatGPT, COPILOT, etc) and text-to-image generators have been used during the writing or editing of manuscripts.

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