

Development and characterization of nano-chitosan-polyvinyl alcohol-glycerol - chlorhexidine films for wound healing applications

Abstract

The process of wound healing is intricate and dynamic, requiring advanced medical care for a speedy recovery. Currently available semi-permeable wound dressings fall short in efficiently absorbing excess wound exudates to facilitate moist wound care, a crucial factor for promoting effective wound healing. Additionally, efforts have been made to enhance the antibacterial properties of wound dressings by incorporating antimicrobials into films. In this study, films were developed using nanochitosan-polyvinyl alcohol-glycerol-chlorhexidine (NC-PVA-GLY-CLX) for applications in wound healing. The dressings were subjected to various characterizations, including UV-Visible spectrometry, FTIR, SEM, TGA, and XRD. Furthermore, the resulting film dressings underwent evaluation for their *in vitro* bactericidal properties. The findings indicated that the incorporation of chlorhexidine (CLX) into the films has positively influenced the fluid holding capacity and the active period of the dressing, creating a favourable environment for effective wound healing.

1. Introduction

The wound healing process is intricate and dynamic pathophysiological process, necessitating advanced medical interventions for effective recovery and well-being [1, 2]. Film-based dressings play a pivotal role in wound care, demonstrating significant potential in treating acute to moderate wounds [3]. Particularly, film-based dressings, known for their superior flexibility, exhibit promising applications, especially in periodontal wound care [4]. However, existing semi-permeable films often fall short in moist wound care [5]. Therefore, the development of wound care devices capable of absorbing exudates is highly favored. Ideal wound dressings should possess crucial attributes such as bacterial permeability, gaseous exchange, sufficient water vapor transmission (WVTR), biocompatibility, and biodegradability [6] for effective healing.

Various natural polymers are extensively explored in the development of wound dressings, incorporating potential cross-linkers and plasticizers to achieve optimal techno-functional attributes suitable for wound dressing applications [7, 8]. Among these polymers, chitosan stands out due to its viscoelastic, thermal, rheological, physical, physico-mechanical, and structural properties. Chitosan, a cationic (positively charged) natural polymer derived from marine shrimps and crab shells, exhibits non-immunogenicity, non-cytotoxicity, antibacterial capabilities [9] biodegradability, and biocompatibility [10], making it optimal for wound healing applications. Chitosan's unique properties arise from the presence of primary amines along its backbone, allowing it to form complexes with anionic counterparts for modulated bio-functionality optimal for wound healing [11]. Polyvinyl alcohol (PVA), when combined with

chitosan, enhances film properties, offering good tensile strength, flexibility, biodegradability, and biocompatibility [12]. Glycerol, a non-volatile polyol, acts as a plasticizer in film-forming solutions to enhance film flexibility [13].

Chlorhexidine (CLX), a gold standard antiseptic agent, exhibits effective antibacterial properties due to its amphipathic symmetrical structure. Its bacteriostatic and bactericidal actions make it a prominent candidate for wound healing applications [14]. Incorporating CLX into chitosan films enhances bactericidal properties and enables sustained release from the polymeric matrix, promoting reduced cytotoxicity for fibroblast cells. The protective barrier provided by chitosan facilitates tissue regeneration, while chlorhexidine prevents infections, expediting the healing process [15]. Nanoparticles synthesized from natural polymers, such as nanochitosan, gain prominence due to their small size and large surface-to-volume ratio [16]. Nanochitosan, produced through ionic gelation, exhibits reproducibility and nanosize particles with a highly positive surface charge [9]. This study explores the synthesis of nano-sized chitosan reinforced with chlorhexidine, plasticized with glycerol, and cross-linked with PVA to create bioactive films for wound healing applications. The synergistic effects of chitosan, PVA, and CLX are comprehensively examined for their functional and bioactive properties through various characterizations, including UV-visible spectroscopy, FTIR, SEM, TEM, TGA, and XRD. The films are further evaluated for their bactericidal properties in a dynamic wound environment.

2. Materials and Methods

2.1. Materials

Chitosan (Low molecular weight) with a deacetylation level of 75% and polyvinyl alcohol (PVA) were acquired from Sigma-Aldrich Chemical Pvt Limited, located in Bangalore, India. Glycerol and all other substances utilized in the current investigation were of either analytical grade (AR) or guaranteed grade (GR). Chlorhexidine base gifted from AB Shetty Dental College, Mangalore.

2.2. Preparation of Nano chitosan

1% chitosan was used to prepare Nanochitosan using sodium tripolyphosphate (STPP), which was added drop by drop into chitosan kept on a magnetic stirrer and solution was stirred for 1 h. The solution was subjected to ultra-sonication for the production of nanochitosan [17].

2.3. Casting of films

Polyvinyl alcohol (PVA) and 0.05% chlorhexidine were introduced to the prepared nanochitosan solution, and the mixture was stirred continuously until a transparent solution was achieved. The fully dissolved solution was then poured onto a tray and left overnight in a hot air oven at 55°C to produce a chlorhexidine nanochitosan film [17].

2.4. Characterisation of Nanochitosan film incorporated with chlorhexidine

2.4.1. Particle size and Zeta potential

The determination of particle size, polydispersity index (PDI), and zeta potential for nanochitosan solutions, nanochitosan with added chlorhexidine, and chlorhexidine (0.05%) was conducted employing dynamic light scattering analyses (DLS). These measurements were performed using a Zeta Sizer Nano Series instrument from Malvern in Worcestershire, UK.

2.4.2. UV VIS spectra

The UV–VIS spectra of nanochitosan solutions, nanochitosan with the integration of chlorhexidine, and chlorhexidine (0.05%) solutions were examined within 200-700 nm wavelength range. This analysis was performed using a double beam spectrophotometer (Shimadzu, UV-1800, Japan). A graph illustrating absorbance against wavelength was generated to obtain the absorption spectra.

2.4.3. FTIR

To identify specific chemical groups in the material Fourier Transform Infrared Spectroscopy (FTIR) was carried out. The FT-IR spectra of the films were inspected using an FTIR spectrometer (Nicolet iS50) across the wavenumber range of 4000 to 500 cm^{-1} , with 64 scans at a resolution of 4 cm^{-1} .

2.4.4. SEM

The surface characteristics of the films were assessed through SEM images captured using a Jeol 6390LA/ OXFORD XMX N, a scanning Electron Microscope equipped with an Energy Dispersive X-Ray Spectroscopy (SEM-EDAX) operating at an acceleration voltage ranging from 0.5 to 30 kV. Various magnifications were employed to capture the surface morphologies.

2.4.5. X- ray diffraction (XRD) pattern

The films were subjected to X-ray powder diffraction analysis using a Bruker D8 Advance diffractometer. Cu $K\alpha$ radiation (40 kV, 80 mA) was employed as the X-ray source. The XRD patterns of the samples were recorded over the 2θ range of 3-10 initially and the final

scanning at 10–80, utilizing a fixed time mode at room temperature and a scanning rate of 4° min⁻¹.

2.4.6. Thermal characteristics

The specimens underwent a temperature increase from 40°C to 750°C at a heating rate of 10°C/min, utilizing the thermogravimetric analyzer TGA-DTA Perkin Elmer STA6000, specifically the Perkin Elmer Diamond model. The analysis involved the examination of the films through DTG (derivative thermo-gravity) and DTA (differential thermal analysis)

2.4.7. Moisture content

Film specimens measuring approximately 2x2 cm were precisely weighed on a dried petri plate and then positioned in a hot air oven (Rotek Instruments, B & C Industries, Cochin, and India). The oven's temperature was maintained at 105 ± 5 °C for a duration of 24 hours. Subsequently, the petri plates were taken out from the hot air oven and allowed to cool within desiccators containing silica gel at room temperature. This process of heating and cooling was iterated until a consistent weight was attained. The determination of the total moisture content was conducted utilizing the equation provided below [18].

Where, W1 = Initial dry weight of the sample; W2 = Final dry weight of the sample

2.4.8. Film solubility

The assessment of film pseudo-exudate fluid solubility was conducted by determining the percentage of dissolved dry matter after immersing film samples. Each film variant (1 cm X 1 cm) underwent an initial drying at 60 °C for 2 hours to establish the initial dry matter (W1). Subsequently, each film was placed in 5 mL of distilled water within a petri plate and stirred gently for 24 hours. After this period, the film specimens were taken out, and their undissolved final dry weight (W2) was determined following 24 hours of drying in an oven set at 105 °C [19]. The solubility (WS) of the sample was calculated using the following formula.

Where, W1 = Initial dry weight of the sample; W2 = Final dry weight of the sample

2.4.9. Swelling Index

The swelling experiments were conducted using film strips measuring 1 × 1 cm, which were immersed in pseudo-exudate fluid (PEF). The PEF composition consisted of 2% bovine serum albumin, 0.02 M calcium chloride, 0.4 M sodium chloride, and 0.08 M trismethylamine in

deionized water, adjusted to pH 7.5. The weight change of the hydrated films was measured at 15-minute intervals for a duration of 150 minutes. After hydration, excess fluids on the film surface were carefully blotted with tissue paper, and the films were promptly weighed using an electronic balance [7].

Where, W_d is dry weight of films; W_s denotes weight of film after swelling

2.4.10. Antibacterial property

2.4.10.1. Diffusion Assay

The initial assessment of antibacterial activity relied on inhibiting bacterial growth on agar plates. This involved placing the membrane over a Mueller Hinton II Agar (Cation-Adjusted) plate coated with a bacterial culture. The plates were prepared, dried, and then spread with a 0.5 McFarland standard concentrated bacterial suspension, as outlined by Visnuvinayagam et al. [20]. After 24 hours, the absence of bacteria clearance beneath the film indicated the presence of antimicrobial activity in the film.

2.4.10.2. Modified Growth Inhibition Assay

To assess the antibacterial efficacy of the film, a modified analysis of growth inhibition was performed. Small film pieces were weighed, and 15 mg of the film was placed into triplicate wells of a 96-well plate. Each well received Mueller Hinton Broth II (Cation-Adjusted) (MHB) (100 μ l), followed by the addition of 10 μ l of bacterial culture (approximately 5×10^5 CFU/mL) to each well, excluding the negative control wells (also in triplicate). In the positive control wells, no film was introduced. Optical density was measured every two hours, and OD values were adjusted by subtracting the initial OD value at 0 hours to eliminate background noise or errors [21].

2.5. Statistical Analysis

The experimentations were conducted in triplicate, and the reported values represent averages with standard deviations. Collected data for various measured parameters were presented and subjected to analysis using the t-test with a significance level at 5% using SAS 9.3

3. Results and Discussion

3.1. Particle size

The characterization of particle size distribution holds paramount significance in determining the functionality of a substance, offering insights into the average particle size.

Table 1 provides a detailed account of the average particle size observed in the film-forming solutions. Particularly in applications such as drug delivery systems, a preference for smaller particle sizes is commonly acknowledged [22]. The Polydispersion Index (PDI) serves as a metric for assessing the distribution or uniformity of particle sizes within a given sample. A lower PDI value signifies a more uniform distribution, indicating a narrower range of particle sizes. The diminutive size of the particles resulting from the chitosan and TPP reaction can be attributed to the utilization of low molecular weight chitosan in this study. Hydrogen bonds developed by the amino and hydroxyl groups of chitosan, in conjunction with the hydroxyl groups and oxygen atoms of water influences the chitosan's charge stability. The strength of the hydrogen bond attraction plays a pivotal role in the propensity of these particles to agglomerate [23]. Additionally, the addition of TPP increases the ionic strength and overall stability of particles being linked by their ions, resulting in aggregate formation [24]. The reduced particle size of NC-CLX is attained through the stabilizing influence of chitosan.

Sl. No.	Sample Name	Particle size (d.nm)	Poly dispersion index (PDI)	Zeta potential (mV)
1.	Chitosan	1989± 2.64	0.35	54.05± 1.00
2.	Nanochitosan (NC)	528.57±2.31	0.40	52.03± 2.00
3.	Chlorhexidine (CLX)	478.77±2.85	0.23	48.87±2.00
4.	Nanochitosan Chlorhexidine (NC- CLX)	319.07± 1.01	0.42	45.20±0.98

Table 1 Particle size, PDI and zeta potential of the formulations

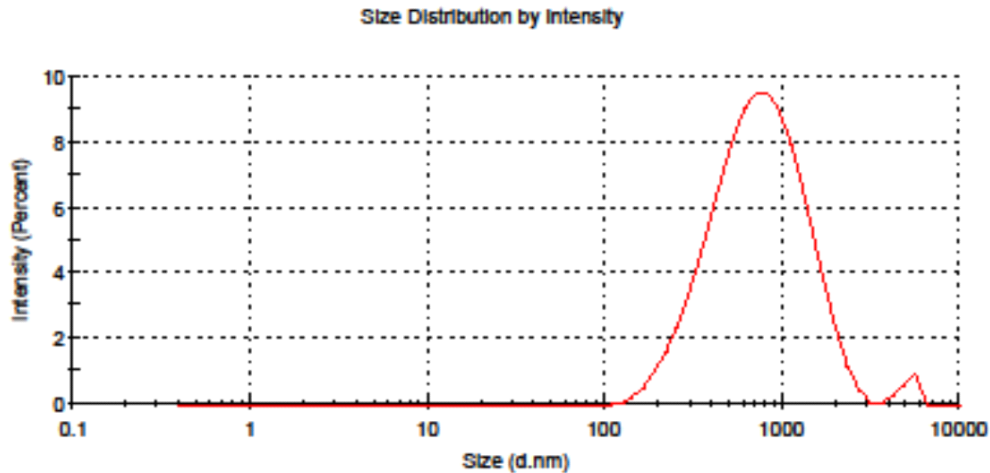


Figure 1. Particle size of Nanochitosan in solution

3.2. Zeta potential

The zeta potential characterizes the potential difference between the dispersal medium and the fluid layer adhering to the charged nanocrystals in a colloidal system. Chitosan, being a polycationic electrolyte with amino groups, imparts a positive zeta potential to the chitosan nanoparticles due to the formation of hydrogen bonds between the amino and hydroxyl groups and the hydroxyl groups or oxygen atoms in water. The presence of TPP decreases the zeta potential of the nanoparticles, as increased crosslinking occurs, leading to the neutralization of protonated amino groups by TPP anions [25]. In the nanochitosan film-forming solution, the cationic chlorhexidine is affixed through TPP, a poly anionic electrolyte with phosphate groups. Upon addition of chlorhexidine, the positive charge on the chitosan particle surface diminishes gradually (as evidenced from the Table 1). This reduction occurs as both intermolecular and intermolecular cross-links form among chlorhexidine, TPP groups, and the amino groups of low molecular weight chitosan [23]. Zeta potentials of the nanoparticles beyond +30 mV or below -30 mV are deemed stable in colloidal suspension systems [26].

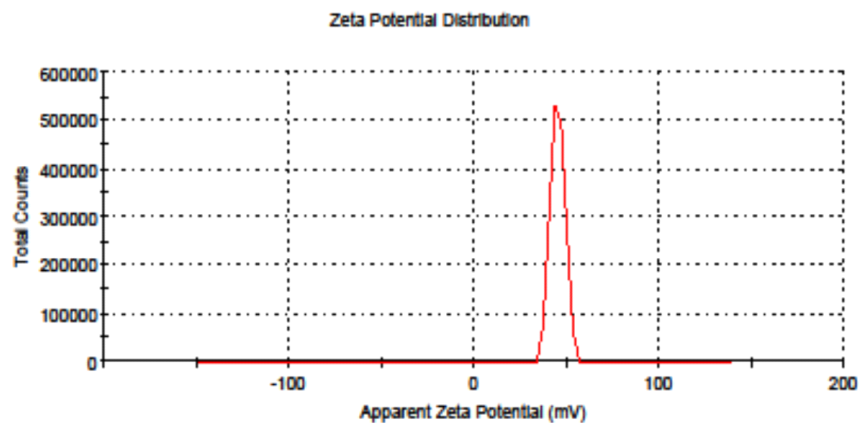


Figure 2. Zeta potential of Nanochitosan formulations

3.3. UV Visible spectra

The utilization of UV-Vis spectra for characterization is favoured due to its straightforward operation and the clarity of results it provides. The UV-Vis absorption spectra, depicted in Figure 3, illustrate nanochitosan, chlorhexidine, and nanochitosan-reinforced chlorhexidine film-forming solutions. Nanochitosan exhibits a broader peak at 289 nm, while chlorhexidine displays a distinct absorption peak at 291 nm. In the nanochitosan-chlorhexidine combination, the chlorhexidine peak at 291 nm merges with the nanochitosan peak, indicating a combination. The introduction of the new compound is evident from the UV spectrum results [27]. The merging of absorption peaks from both chitosan and chlorhexidine in the UV-Visible

spectrum and the shift broadening at 291 nm suggests the successful incorporation of chlorhexidine into the chitosan solution. This indicates a proper combination of the two components in the reaction system [28].

Figure 3. UV Visible spectra of formulations

3.4. FTIR

Fourier transform infrared spectroscopic (FTIR) analysis imparts insights into the functional properties, establishing correlations between functional groups and the structure of the blended film. The spectrum reveals the characteristic absorption band of nanochitosan at 3270 cm^{-1} , indicative of the $-\text{NH}$ group in chitosan, widened by physical interactions with TPP. Strong peaks at 2939 cm^{-1} , 1646 cm^{-1} , 1561 cm^{-1} , 1328 cm^{-1} , and 1037 cm^{-1} signify asymmetrical and symmetrical stretching in CH_2 groups, NH_3^+ stretching, $\text{C}=\text{O}$ stretching in amides, NH bending, and OH in-plane bending in alcohols, and $\text{P}=\text{O}$ stretching, consistent with the findings of Vijayalakshmi et al. [29]. The peak at 1142 cm^{-1} indicated overlapping of $\text{C}-\text{O}$ stretching in polysaccharide and the formation of chitosan nanoparticles due to interactions between ammonium and phosphate ions in chitosan nanoparticles [29]. In the FTIR spectrum of chlorhexidine (CLX), characteristic bands at 1640 cm^{-1} signify the double bond $\text{C}=\text{N}-\text{H}$ stretching vibration, while bands at approximately 1532 cm^{-1} and 1400 cm^{-1} correspond to aromatic $\text{C}=\text{C}$ stretching modes. Peaks at 2923 cm^{-1} and 2864 cm^{-1} attribute to CH_2CH stretching in methylene groups of CLX [15]. Overlapping of absorption bands between chitosan and chlorhexidine occurs, resulting in increased intensity, notably in the $3750\text{--}2750\text{ cm}^{-1}$ region [30].

Figure 4. FTIR Spectrum of Chlorhexidine, Nanochitosan and nanochitosan chlorhexidine

3.5. SEM

Scanning electron microscopy (SEM) was conducted to elucidate the intricate surface topography of the film specimens. Micrographs of nanochitosan films (Fig 5. a) unveiled the presence of rod-like structures. The film surfaces exhibited a notable smoothness and homogeneity, devoid of pores, cracks, or air bubbles, a consequence of the meticulous drying process (solvent casting). The conspicuous absence of voids or cracks in the film points to the application of optimal temperatures during film preparation ($< 80^\circ\text{C}$). Noteworthy is the absence of agglomerations in nanochitosan films, affirming the uniform dispersion of component particles, as highlighted in the study by Zolfi et al. [31]. The findings align with prior research, notably the rod-like structures documented by Vijayalakshmi et al. [29]. Upon the introduction of

chlorhexidine into the films, a discernible alteration in surface appearance manifested with the emergence of flake-like structures interspersed among the nano-rods. These distinct rod-like nanostructures offer a nuanced avenue for bespoke applications, particularly in sustained drug delivery scenarios where the elongated shape may confer advantages in terms of enhanced penetration or interaction with target cells.

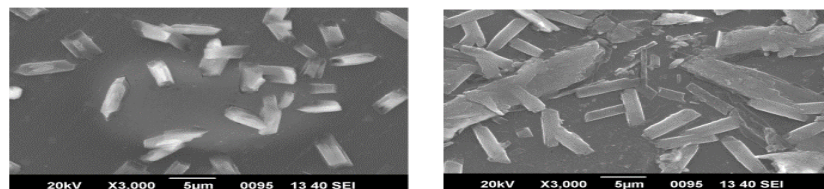


Figure 5a. Nanochitosan and 5b. Nanochitosan Chlorhexidine film

3.6. XRD

The application of X-ray diffraction analysis served as a non-invasive analytical method for discerning modifications in the crystalline attributes, providing profound insights into crystalline phase identification, chemical composition, crystallographic structure and physical properties of materials. Consequently, a meticulous examination of the X-ray diffraction patterns of the fabricated membranes was undertaken, acknowledging the pivotal role played by the crystalline-to-amorphous ratio in comprehending the membrane development during gelation, as underscored by Vinodhini et al. [32].

The X-ray diffraction pattern of the nanochitosan film, depicted in Figure 6 displayed a diffractogram characterized by a broad peak around $2\theta=19.6^\circ$. This broadening of peaks is a result of the deformation of crystalline regions due to the heightened packing of chitosan chains through ionic crosslinking. While the original chitosan molecules exhibited a certain regularity conducive to the formation of crystalline regions, crosslinking with tripolyphosphate led to a diminished crystallinity of chitosan. This reduction is ascribed to the distortion of hydrogen bonds in the original chitosan caused by the substitution of hydroxyl and amino groups, effectively disrupting the orderly packing of chitosan chains and resulting in the formation of amorphous nanochitosan. Furthermore, the absence of prominent peaks at $2\theta= 8.4$ and 8.6 (attributable to a dispersion effect) signifies the utilization of a low concentration of chlorhexidine (CLX) ($< 0.5\%$) in the films [29]. The presence of TPP in polymeric blends is indicated by the $2\theta = 34.26$ peak [24].

Figure 6: Nanochitosan Chlorhexidine Film

3.7. TGA

Thermogravimetric analysis was employed to explore the impact of temperature variations on the mass loss of the films, providing reliable insights into polymer degradation concerning temperature, decomposition temperature, and residual content under a nitrogen atmosphere, as documented by Ma et al. [33]. In the case of the nanochitosan film, degradation and weight loss were observed at rates of 3.057%/min at 185.32 °C, 4.068%/min at 270.68 °C, 2.5%/min at 380.43 °C, 4.8%/min at 480.42°C, and 2.1%/min at 480.42°C. The primary degradation transpired within the temperature range of 184–400 °C, with the maximum weight loss rate occurring at 270 °C. Notably, Meera et al.[34] indicated that the decomposition of the principal polymeric chains of chitosan commences at 160°C, and beyond 450 °C, thermal degradation is attributed to the cross-linkages in nanochitosan [35].The modification of the nanochitosan film with the cationic drug chlorhexidine resulted in an increased number of decomposition steps. The initial degradation commenced at 74.13°C with a decomposition rate of 1.866%/min, possibly attributable to film dehydration or the removal of residual solvent [36]. The maximum weight loss occurred between 160-400°C, with decomposition rates of 3.06%/min at 229.54°C, 5.97 %/min at 270.44°C, and 2.19 %/min at 376.78°C. Additional thermal degradation events were observed at rates of 3.02 %/min at 499.48°C and 3.24%/min at 558.37°C, indicating the thermal stability of the resultant films.

Figure 7. TGA of Nanochitosan Film

Figure 8. TGA of Nanochitosan Chlorhexidine Film

3.8. Functional properties

3.8.1. Moisture content (MC)

In the realm of biomedical applications for polymeric films, a pivotal aspect that demands meticulous consideration is the moisture content. Its significance is particularly pronounced in wound healing applications, exerting direct influence on the microenvironment of the wound and

intricately shaping various physiological processes essential for effective wound healing [37]. Although both chitosan and PVA possess hydrophilic components, their distinct affinities and interaction capacities with water are noteworthy. The moisture content assumes the role of a modifier, significantly impacting the mechanical properties of the films [38]. According to the findings, nanochitosan films exhibit a higher moisture content, registering at $35.33\pm 2.67\%$ and the results were in agreement with the study of Chopra et al., [39]. A substantial decrease to $22.44\pm 2.03\%$ is evident in the moisture content of the film upon the addition of chlorhexidine, an amphipathic molecule featuring both hydrophilic and hydrophobic groups [40]. Maintaining an optimal level of moisture in the wound bed, rather than fostering a dry environment, proves conducive to wound healing. This approach prevents tissue dehydration and subsequent cell death, fosters angiogenesis, and facilitates the removal of dead tissue, thereby promoting an environment conducive to effective wound healing [41].

3.8.2. Film solubility

Film dressings, characterized as thin transparent polymeric sheets, create a conducive environment for moist wound healing [42]. Notably, a substantial decrease in solubility was evident in pseudo exudate fluid during film solubility studies, indicating resilience to fluids. This reduction in film solubility signifies the stability and resistance of the formulation to water [43]. The incorporation of chlorhexidine led to a significant reduction in solubility in pseudo exudate fluid, decreasing from 25.27 ± 0.70 to 12.96 ± 0.73 consistent with these studies by Abdul-Rahman and Abas [44] reveal a similar decrease in solubility values for chitosan film upon the addition of supporting substances. This trend may be attributed to the robust interaction between phenolic compounds and the polysaccharide chain of chitosan, potentially diminishing the availability of amine and hydroxyl groups in the polymer for interaction with water. Consequently, the water solubility of the film decreases. In the realm of drug delivery and pharmaceutical applications, films with reduced solubility can be tailored to release active ingredients gradually and predictably. This controlled release mechanism ensures a consistent and sustained therapeutic effect, thereby enhancing patient compliance and treatment efficacy.

Functional tests	NC	NC-CLX
Moisture Content (MC) %	35.33 ± 2.67	22.44 ± 2.03
Film Solubility (%)	25.27 ± 0.70^a	12.96 ± 0.73^b

^a and ^b indicates significance level at 5% using Independent t-test

Table2. Moisture content and film solubility of the films

3.8.3. Swelling Index (Fluid handling capacity)

Swelling plays a crucial role in the evaluation of biodegradable films, influencing their water resistance, particularly in humid conditions [45]. The permeation of solvent into the chitosan polymer matrix induces the expansion of the composite membrane and causes localized relaxation of segments within the polymer matrix [18] ascertained by ionic repulsion mechanism. The higher swelling index of NC-CLX films indicates its fluid handling capacity at pH 7.4 (acute wound), which remain optimal for wound healing [45].

3.9. Anti-bacterial activity

In this investigation, two types of nanochitosan films were employed—nanochitosan (NC) and nanochitosan chlorhexidine (NC-CLX). Significantly, films incorporating chlorhexidine displayed a discernible inhibition zone. Using a 0.5 McFarland Standard concentration of *Methicillin-Resistant Staphylococcus aureus* (MRSA) culture spread on Mannitol Salt agar (MSA), the films were placed over the bacterial culture. The observed inhibition zone hierarchy was NC-CLX > NC. The antimicrobial efficacy of nanochitosan, characterized by NH_3^+ groups, involves interaction with the negatively charged surface of microbial cell membranes. This interaction disrupts cell functions by breaking components or inhibiting cell activity, resulting in microbial cell death. Nano chitosan's capability to penetrate microbial cell membranes enables it to bind to DNA, hindering the synthesis of RNA, enzymes, and proteins [46]. Conversely, cationic chlorhexidine binds to the negatively charged cell wall of bacteria, attacking cell membranes and causing cell death [47]. The shift from traditional assays to micro-well-based assays is attributed to their simplicity, cost-effectiveness, and ability to analyse a larger sample volume in a shorter timeframe

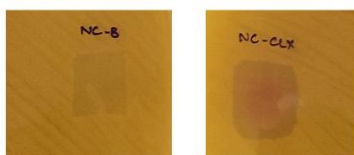


Figure 9. Anti-bacterial activity NC and NC-CLX film

In this study, Modified growth inhibition assay (MGIA) was applied, utilizing 96-well microtiter plates to evaluate the active period of films concerning its applications to the wound. Both *Escherichia coli* and MRSA (wound associated bacteria) were cultured in the 96-well plates for a period of 10 hours. Results obtained from the MGIA indicated that NC-CLX

demonstrated a more robust antimicrobial activity compared to NC. The presence of the film impeded bacterial growth, evident from the reduction in optical density measurements. This inhibitory effect on bacterial growth was further supported by the sustained growth observed in the positive control wells, where only bacteria were present without the film. Conversely, the negative control wells, lacking both bacteria and the film, exhibited unaltered OD values. The findings of this study suggest that the NC-CLX film possesses superior antimicrobial properties. The use of MGIA in 96-well microtiter plates provided a convenient and effective method for evaluating the antimicrobial characteristics of the films. In accordance to the results, the NC-CLX films offered an active period of > 10 hrs. at experimental conditions.

Figure 10. Modified growth inhibition assays for the NC and NC-CLX Films of E. coli

Figure 11. Modified growth inhibition assays for the NC and NC-CLX Films of MRSA

4. Conclusion

In this investigation, nanochitosan composite films infused with chlorhexidine were formulated with the primary goal of amplifying their antimicrobial efficacy for wound healing applications. The intrinsic antibacterial attributes of chitosan, stemming from its primary amine groups, were fine-tuned through the incorporation of chlorhexidine. The films underwent thorough physiochemical analyses, encompassing XRD, SEM, FTIR, and TGA, to elucidate their characteristics. The amalgamation of chlorhexidine into nanochitosan films not only broadens their functional spectrum, including enhanced fluid handling capacity and resistance, but also bestows antimicrobial advantages, biocompatibility, and the potential for multifunctional properties. This unique combination positions these films as a highly promising option for wound care and diverse therapeutic applications in the medical and healthcare domains.

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