

Synthesis and characterization of nickel(II), cobalt(II) and Zinc(II) complexes with 5-chloropyridine-2-carbaldehydethiosemicarbazone

Abstract

5-chloropyridine-2-carbaldehydethiosemicarbazone (HL1) and its nickel(II), cobalt(II), Zinc(II) and copper(II) complexes: $[\text{Ni}(\text{HL1})_2]$, $[\text{Co}(\text{HL1})_2]$, $[\text{Zn}(\text{HL1})_2]$, $[\text{Cu}(\text{HL1})_2]$, $[\text{Cu}(\text{HL1})]$, were synthesized and characterized. The ligand has been characterized by elemental analyses, IR, ^1H NMR and Mass spectroscopy. The dentate nature of the ligand is evident from the spectral obtained data and literature.

Keywords: Synthesis and characterization, thiosemicarbazone, Ni complex, Cu and zincic complex, 5-chloro 2-furaldehyde

1. Introduction

Heterocyclic thiosemicarbazone and their metal complexes are among the most widely studied compounds of thiosemicarbazone that containing important hetero atoms like sulphur and nitrogen. Heterocyclic thiosemicarbazone possess Structural simplicity, easy synthetic procedure, shorter course of reaction and wide pharmacological profile make them a class extensively studied and explored. They offer a wide spectrum of pharmacological activities like analgesic and anti-inflammatory[1], anti-viral[2], antibacterial[3], anticonvulsant[4], anti-HIV[5], anti leishmaniac[6], antifungal[7], anti-malaria[8], anticancer[9], neurotropic[10], anti trypanosomal[11], antitubercular[12], insulin like[13], and human and rat cholinesterase[14] activities. Metal complexes of thiosemicarbazones have also been found to be useful in radiopharmacy for therapeutic as well as diagnostic purposes[15]. Anti-cancer properties of these derivatives have been studied widely. Recently, there have been frequent studies of metal complexes of heterocyclic thiosemicarbazones[16] Many of these have featured either 2-formylpyridine[17-20] or 2-acetylpyridine[21-24] thiosemicarbazones; of 2-benzoylpyridine thiosemicarbazone[25]. To our knowledge, studies of the metal complexes of the of chloride substituents at C5 carbon of pyridine-2-carbaldehydethiosemicarbazone and its complexes was not investigated enough. The present communication reports the synthesis, characterization of Ni(II), Co(II) and Cu(II) complexes with 5-chloropyridine-2-carbaldehydethiosemicarbazone.

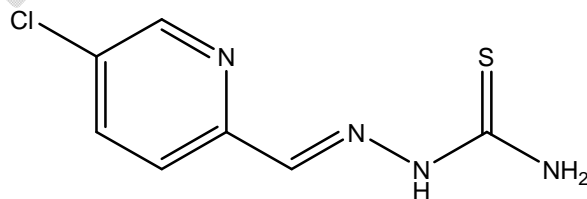


Fig (1): The structure of 5-Chloro-2-pyridinecarboxaldehyde thiosemicarbazone((E)-2-((5-chloropyridin-2-yl)methylene)hydrazine-1-carbothioamide

2. Materials and Methods

2.1. Materials

5-chloropyridine-2-carbaldehyde (Merck) and thiosemicarbazide (Aldrich) were used as received. The metal salts $\text{CoCl}_2\cdot\text{H}_2\text{O}$, $\text{NiCl}_2\cdot 4\text{H}_2\text{O}$, $\text{CuCl}_2\cdot 2\text{H}_2\text{O}$ (Merck) were used as supplied. Solvents used for the reactions were used as received. The ligand has been prepared by the method reported earlier [13]. As an example some characteristics of the ligand are summarized below.

2.2 Preparation of the ligand (HL1):

5-Chloro-2-pyridinecarboxaldehyde (0.283 g, 0.002mol) in ethanol (10 cm^3) was added to thiosemicarbazide (0.182 g, 0.002 mol) in ethanol-water solution (V/V 1 : 1, 80 cm^3) followed by the addition of drops of glacial acetic acid. The reaction mixture was refluxed for 3 hours. The obtained precipitate (5-Chloropyridine-2-carbaldehyde thiosemicarbazone) (HL1) was filtered off, washed with water and ethanol and recrystallized from ethanol, and then was dried in vacuum.

2.3 Preparation of the Complexes

General procedure for the preparation of the metal complexes (1–5)

Complexes 1–5 were prepared by direct reaction between the ligand and the corresponding metal salts

2.3.1 Synthesis of [Co (HL1)2]

To a solution of Cobalt (II) Chloride $\text{CoCl}_2\cdot 6\text{H}_2\text{O}$ (0.0005mol, 0.1189 g) in 5ml of methanol was added a solution of ligand 5-Chloropyridine-2-carbaldehyde thiosemicarbazone (0.001mol, 0.215g) in 25 ml of methanol. The mixture was refluxed for 2 hours (70-80 oC) and then stirred for 12 hours then filtered. The volume of filtrate was left for slow evaporation at room temperature. After two days large of dark brown colored crystals was obtained.

2.3.2 Synthesis of [Ni (HL1)2]

The hexahydrated nickel chloride, $\text{NiCl}_2\cdot 6\text{H}_2\text{O}$ (0.0005 mmol, 0.119 g), was dissolved in 30ml distilled water. An ethanolic solution of 5-Chloropyridine-2-carbaldehyde thiosemicarbazone (0.001mmol; 0.215 g, 10ml) was added slowly while stirring. The mixture was refluxed for 3 h. After cooling at room temperature, a green precipitate appeared. It was filtered, washed with small amounts of absolute ethanol and finally dried in vacuum over silicagel(light brown) . All the compounds were washed and dried in the same way.

2.3.3 Synthesis of [Zn(HL1)2]

0.136g, 0.001mol ZnCl_2 was dissolved in 20 ml warm absolute ethanol with stirring. To a solution containing 0.429g, 0.002 mol 5-Chloro-2-pyridine -2-carboxaldehydethiosemicarbazone dissolved in 20 ml of absolute ethanol drop wise of NaOH

(1M) was added. The mixture was refluxed for two hours. The separated brown powder was filtered washed with ethanol and to dry in air.

2.3.4 Synthesis of [Cu (HL1)2]

A quantity of 5-Chloro-2-pyridine -2-carboxaldehydethiosemicarbazone (0.430g, 0.002 mol) was dissolved in warm 20 mL ethanol with stirring and to this solution was added slowly a solution of (0.134 g 0.001mol) of the metal salts CuCl₂.2H₂O dissolved in 20 mL ethanol, then added drops of NaOH (1M). Although the complex appeared instantaneously the mixture was reflux on water bath for two hours. The separated green solid was left to dry in air after filtered and washed with ethanol.

2.3.5 Synthesis of [Cu(HL1)]

A quantity of 5-Chloro-2-pyridine -2-carboxaldehydethiosemicarbazone (0.215g, 0.001 mol) was dissolved in warm 20 mL ethanol with stirring and to this solution was added slowly a solution of (0.134 g 0.001mol) of the metal salts CuCl₂.2H₂O dissolved in 20 mL ethanol, then added drops of NaOH (1M). Although the complex appeared instantaneously the mixture was reflux on water bath for two hours. The separated green solid was left to dry in air after filtered and washed with ethanol.

2.4 Physical measurements

- **Elemental analyses** for (C, H, N and S) were performed using a Heraeus Carlo Erba 1108 elemental analyzer.

-**Melting points** were determined with a digital melting point apparatus using capillary technique.

-**Mass spectra** were recorded with a Micromass LCT electrospray (Qtof Micro YA263) mass spectrometer.

-**NMR spectra** were recorded in DMSO-d₆ solution on a Bruker Avance DPX 300 NMR spectrometer.

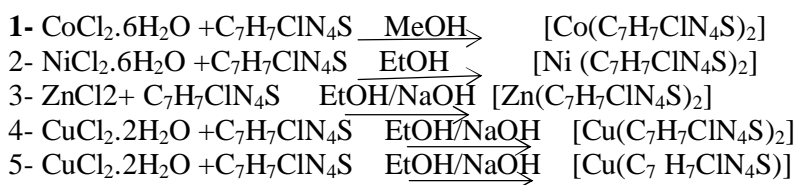
-**IR spectra** were obtained on a Perkin Elmer Spectrum Two IR spectrometer with samples prepared as KBr pellets.

3. Results and discussion

3.1. Synthesis and Physical Properties of the Complexes.

Firstly the ligands(HL1) was prepared from the reaction between 5-Chloro-2-pyridinecarboxaldehyde and thiosemicarbazide in alcoholic medium followed by the addition of drop of glacial acetic acid, the melting point of (HL1) = 201⁰C. The elemental analytical calculation for the ligand was found C, 39.25, H, 3.26; N, 26.1; S, 17.0%, it's quite agreed with literature(C, 39.13, H, 3.25; N, 26.9; S, 17.65%).

After that the preparative reactions for the complexes can be represented by the following equations:



Note that all these reaction needed areflux for certain time as metioned in the above procedure(2.3.1-2.3.5).

All the complexes are microcrystalline or amorphous powder, stable in the normal laboratory atmosphere, and slightly soluble in common organic solvent but completely soluble in DMF and DMSO.

The results of color, M.p and elemental analysis of the ligand 5-chloropyridine-2-carbaldehydethiosemicarbazone alone and its Ni, Co, Cu and Zn complexes were shown in table(1) below

Table (1) Analytical data

Compound	Color	M.p(°c)	Elemental analysis found (Cal)			
			C (%)	H (%)	N (%)	S (%)
5CIPTSC	Brown	154	39.13(39.25)	3.25(3.26)	26.9(26.10)	17.65(15.02)
[Ni (C ₇ H ₇ ClN ₄ S) ₂]	green	226	28.66(27.7)	2.83(2.3)	18.83(18.42)	7.81(10.5)
[Co (C ₇ H ₇ ClN ₄ S)]	Blue	244	30.50(30.16)	2.63(2.51)	19.65(20.11)	8.03(11.49)
[Cu (C ₇ H ₇ ClN ₄ S) ₂]	Dark green	237	30.74(29.86)	2.63(2.48)	20.8(19.90)	11.5(11.37)
[Zn (C ₇ H ₇ ClN ₄ S)]	Brown	216	23.99(23.95)	2.36(2.00)	16.71(16.00)	6.41(7.90)
[Cu (C ₇ H ₇ ClN ₄ S)]	Dark green	233	23.61(24.00)	2.09(2.00)	15.7(16.04)	10.13(9.00)

3.1. IR Spectra

The IR spectral data, table (2) 5-Chloro-2-pyridinecarboxaldehyde thiosemicarbazone, shows bands at 3419cm⁻¹ which is assigned to ν (NH₂) sym.str.vib. The band at 1606 cm⁻¹ is assigned to ν (C=N) sym.str.vib, and ν (C=S) sym.str.vib shows band at 825 cm⁻¹. The band at 3157 is assigned to ν (NH) sym.str.vib .

The Infra-red spectral data of Ni (II) complex with the ligand HL¹ is shown in table (2). The band at 1606 cm⁻¹ assigned to ν (C=N) in the free ligand is shifted to 1616 cm⁻¹ in the spectrum of the complex. This indicates complexation of ligand through the azomethine nitrogen to the Ni (II) ion. The band at 825 cm⁻¹ assigned to ν (C=S) in the free ligand is shifted to 833 cm⁻¹. This suggests coordination of the ligand via the ν (C=S) sulfur.

The Infra-red spectral data of cobalt(II) complex with the ligand is shown in table (2). The band at 1606 cm⁻¹ assigned to ν (C=N) in the free ligand is shifted to higher frequency, (1618 cm⁻¹), in the spectrum of the complex. This indicates complexation of the ligand through the azomethine nitrogen to the cobalt (II). The band at 825 cm⁻¹ assigned to ν (C=S) in the free ligand is shifted to 877 cm⁻¹. This suggest coordination of the ligand via the ν (C=S) sulfur.

The Infra-red spectral data of zinc(II) complex with the ligand is shown in table (2). The band

at 1587 cm⁻¹ assigned to $\nu(\text{C}=\text{N})$ in the free ligand is shifted to 1606 cm⁻¹ in the spectrum of the complex. This indicates complexation of the ligand through the azomethine nitrogen to the zinc(II) ion. The band at 838 cm⁻¹ assigned to $\nu(\text{C}=\text{S})$ in the free ligand is disappeared in the complex spectrum. This may be due to deprotonation of the thiol sulfur atom.

The copper ligands complexes IR spectral data can be explain in the same way which mentioned above for Ni, Co and Zn complexes with slightly different in the position of absorption band between the two Cu complexes as its clear from table (2) the different band in Cu(HL)₂ complexes were shifted to lower frequency in Cu(HL) and this is attributed to lowering of the force constant and dipole moment in the second complexes.

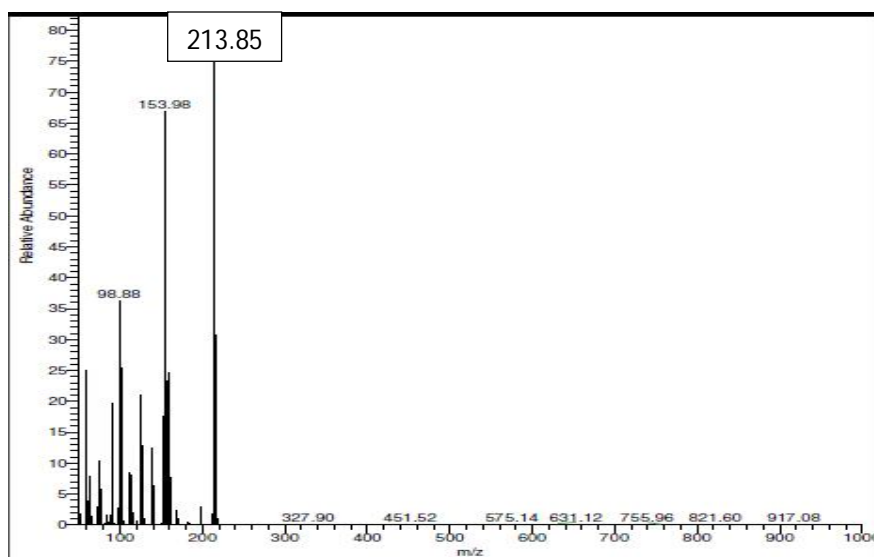
Table (2) Main IR spectral vibrations (cm⁻¹) of the ligand 5CIPTSC and its Ni, Co, Cu and Zn complexes

Compound	$\nu(\text{NH}_2)$	$\nu(\text{C}=\text{S})$	$\nu(\text{C}=\text{N})$	$\nu(\text{M}-\text{N})$	$\nu(\text{M}-\text{S})$
5CIPTSC	3419	825	1606	511	411
[Ni (C ₇ H ₇ ClN ₄ S) ₂]	3267	833	1616	453	432
[Co (C ₇ H ₇ ClN ₄ S) ₂]	3269	877	16018	531	430
[Cu (C ₇ H ₇ ClN ₄ S) ₂]	3461	784	1595	491	441
[Zn (C ₇ H ₇ ClN ₄ S)]	3276	705	1631	522	422
[Cu (C ₇ H ₇ ClN ₄ S)]	3361	686	1625	426	407

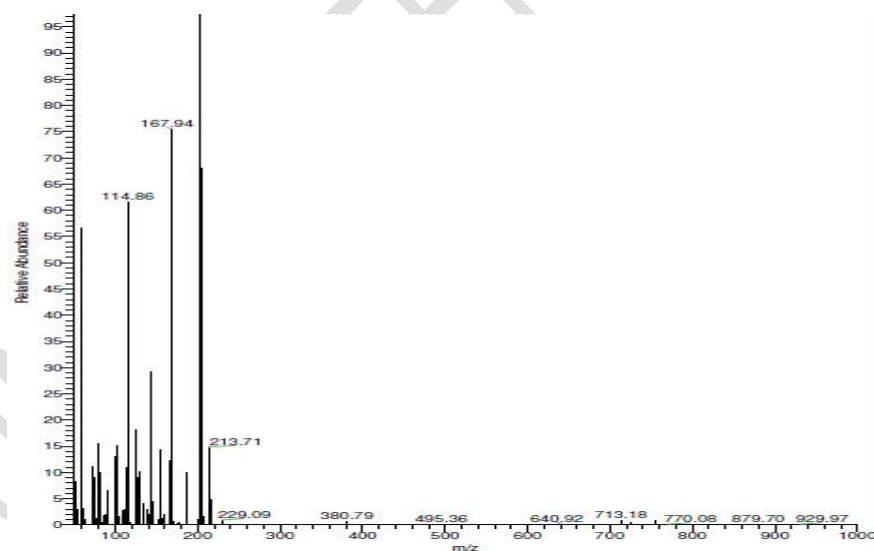
3.2 NMR and Mass Spectra

The ¹H NMR spectrum of (5CIPTSC), shows a singlet at 11.3 ppm due to the NH group next to C=S, while the signal of the proton on C=N double bond appears at 7.50 ppm. It is interesting to note the presence of two broad singlets for the two NH₂ protons, respectively at 7.78 and 8.15 ppm; it means that the free rotation around the C=N bond is blocked because of its partial double bond character [10]. The spectra of the complexes [CoCl₂(5CIPTSC)₂] and [CuCl₂(5CIPTSC)] show no differences as compared to that of the free ligand. This enhances the hypothesis

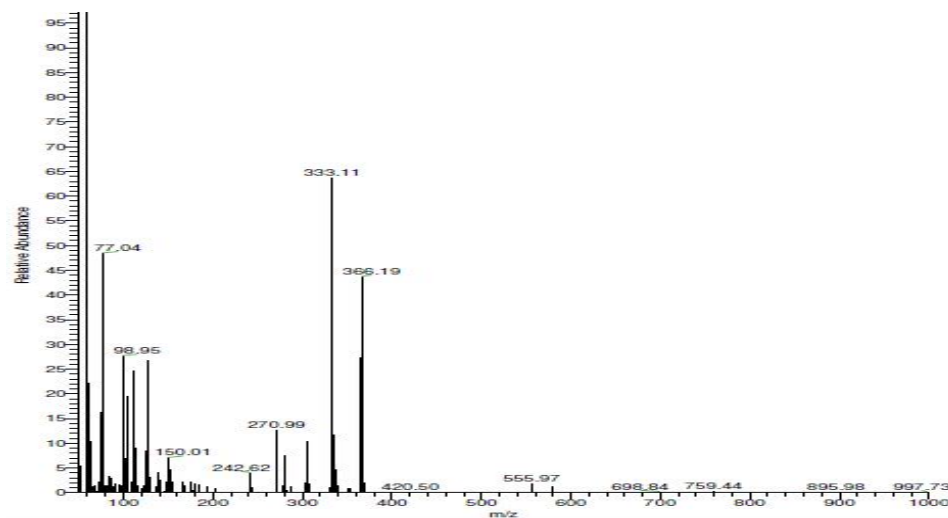
Figs (3) to (6) showed the mass spectrum of the ligand and metal (Co, Ni,Zn and Cu) complexes all the expected fragmentations and the molecular ion peak and m/z peaks appeared successfully



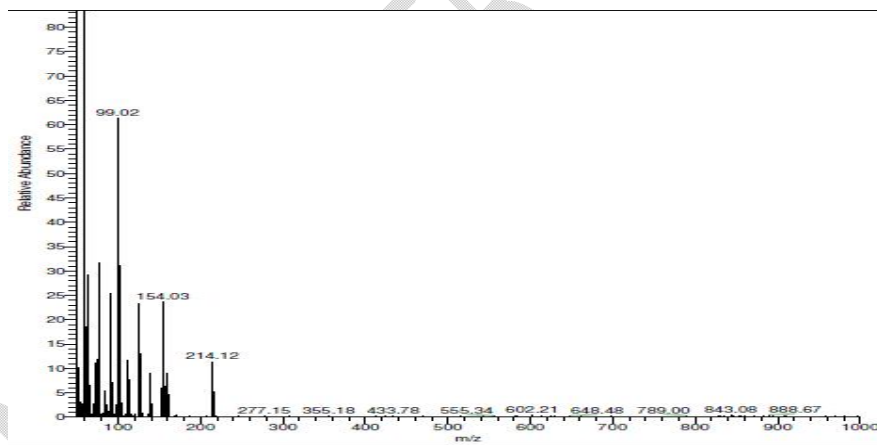
Fig(3) Mass spectrum of 5-Chloro-2-pyridinecarboxaldehyde thiosemicarbazone



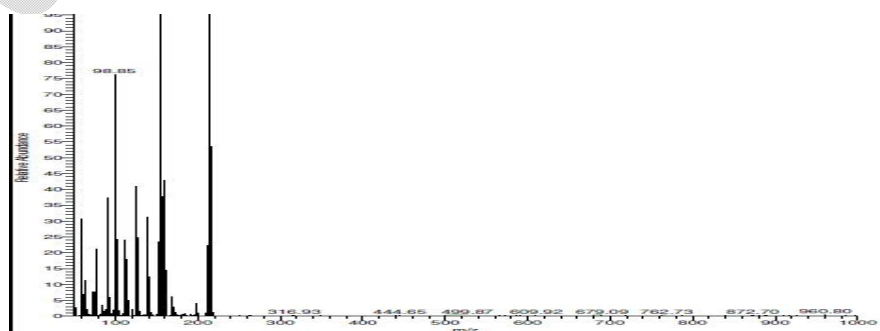
Fig(4) Mass spectrum of [Ni(HL1)₂]



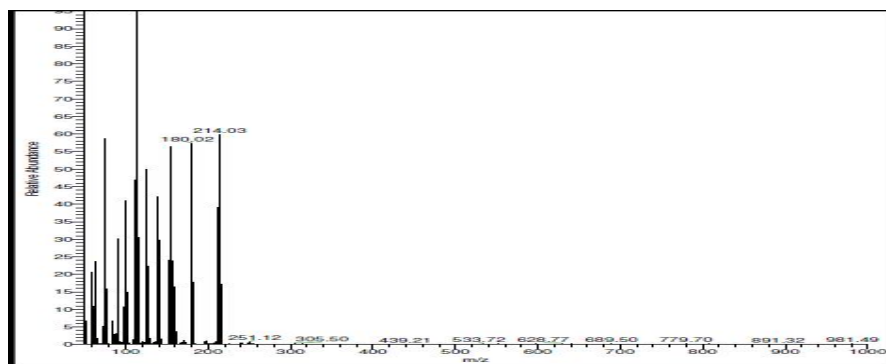
Fig(5) Mass spectrum of [Co(HL1)₂]



Fig(6) Mass spectrum of [Zn(HL1)₂]



Fig(7) Mass spectrum of [Cu(HL1)]



Fig(8) Mass spectrum of [Cu(HL1)₂]

The proposed structure for the different complexes and after our previous discussion of the spectral data obtained for the ligands and its metal complexes were shown below in Figs(9-12) This may suggest a tetrahedral geometry for the complexes (Suni.V et al., 2007).

In the complex of cobalt(II) with the ligand 5-Chloro-2-furaldehyde thiosemicarbazone (HL²), the band at 1587 cm⁻¹ assigned to ν (C=N) is shifted to 1622 cm⁻¹, indicating coordination of the ligand through the azomethine nitrogen. The band assigned to ν (C=S) is shifted from 838 cm⁻¹ to 831 cm⁻¹ indicating the involvement of sulfur in complexation.

The [Cu (HL)₂]₂Cl₂ complex shows bands at 46082 cm⁻¹ (217 nm), 28169 cm⁻¹ (255 nm) and 13850 cm⁻¹ (534 nm). The observed magnetic moment is 1.86 BM, this suggest distorted square planar geometry of the copper center.(Amna Qasem Ali et al ; 2014). For Zn the complex is square planar from the obtained spectral data (Mihail Revenco, et al; 2014).

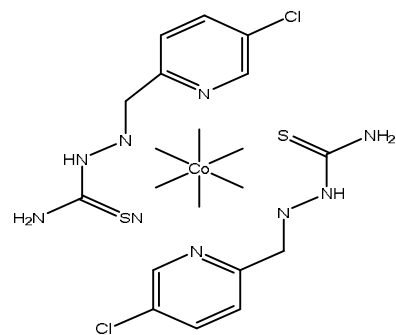


Fig:(9): The probable structure of complex [Co (L1)2]

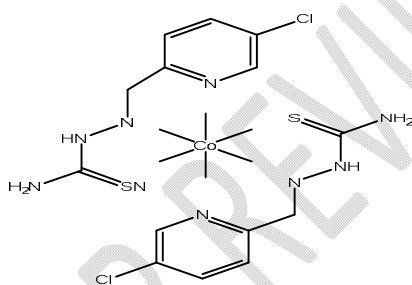


Fig:(10): The probable structure of complex [Ni (L1)2]

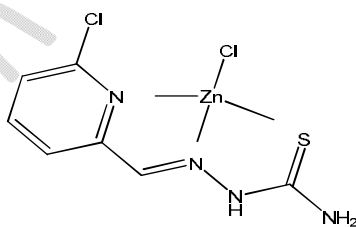


Fig:(11): The probable structure of complex [Zn (L1)]

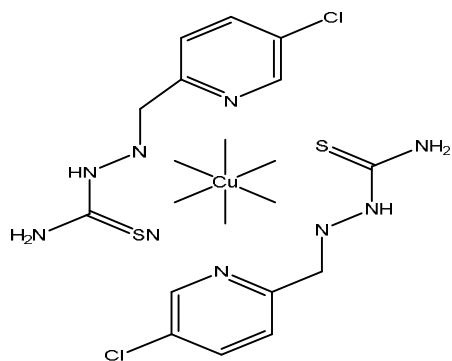


Fig:(12): The probable structure of complex [Cu (L1)2]

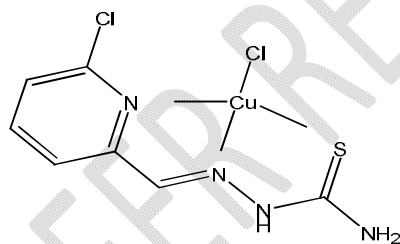


Fig:(13): The probable structure of complex [Cu (L1)]

5. Conclusions

Synthesis and characterization of metal ligand complexes was very important field because these compounds found wide applications, in this study the synthesis and characterization of nickel(II), cobalt(II) and Zinc(II) complexes with 5-chloropyridine-2-carbaldehydethiosemicarbazone had been studied. The complexes were synthesized successfully and the structure of them were indicated from the obtained spectral data, the research goes further in antimicrobial investigation which did not shown here because the results were not complete till now and we will published it in our next coming paper.

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