

**Original Research Article**

**ISOLATION, CHARACTERISATION AND CORROSION INHIBITION EVALUATION OF 1, 3, 5-TRIMETHOYMETHYLBENZENE FROM SEEDS EXTRACT OF *Cola nitida* ON MILD STEEL SURFACE IN HYHROCHLORIC ACID SOLUTION**

**ABSTRACT**

Corrosion inhibition of Mild steel in 1 M hydrochloric acid solution by 1, 3, 5-trimethoymethylbenzene (TMMB) isolated from seeds extract of *Cola nitida* and characterized was studied using weight loss, potentiodynamic polarization and electrochemical impedance spectroscopic methods. The results recorded show that TMMB inhibits corrosion of mild steel in hydrochloric acid solution. The inhibition efficiency by the weight method increased with increase in TMMB concentration, but decreased with increase in temperature. The inhibition efficiencies by the potentiodynamic polarization and electrochemical spectroscopic methods also increased with increase in the concentration of TMMB. The calculated thermodynamic parameters showed that the corrosion inhibition process was endothermic and spontaneous in nature. Potentiodynamic polarization measurements confirmed that TMMB is a mixed – type inhibitor considering the value of  $E_{corr}$  being less than the threshold value of 85 mV. Impedance measurement revealed that the charge transfer resistance offered by TMMB to the migration of aggressive ions on the surface of mild steel in acidic solution was enough to impede corrosion. The best line of fit to the adsorption of TMMB onto mild steel surface was found in Temkin isotherm. Formation of protective dense film on the steel surface in the presence of TMMB compared to the blank was revealed by scanning electron microscopy. The molecular of structure of TMMB was optimized using Gaussian 09 with density functional theory [DFT/B3LYP/6-31G (d,p)]. Physical adsorption has been proposed for the adsorption of TMMB onto mild steel surface.

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**Comment [h3]:** Worked as

**Comment [h4]:** What the author mean by the threshold value

**Keywords:** *Cola nitida*, Mild steel, Acidic solution, Weight loss, Potentiodynamic polarization, Impedance spectroscopy, Density functional theory, Temkin isotherm.

**1 INTRODUCTION**

Materials for construction and some components of oil and gas industries made of mild steel rods, sheets and pipes are often exposed to industrial processes involving acidic solutions. This paves the way for industrial processes such as pickling, chemical and electrochemical etching and industrial acid descaling. The metallic components are susceptible to corrosive actions, which are retrogressive and interfacial reactions of a material (metal, ceramic and polymer) with its environment resulting in consumption of the material [1] and they deteriorate on continued usage.

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The use of inhibitors in corrosion prevention and control in aggressive environments is the most cost-effective. Most of the efficient inhibitors are organic compounds that contain heteroatoms such as nitrogen (N), sulphur (S), oxygen (O) atoms,  $\pi$ -electrons and suitable functional groups. These atoms are found to have higher basicity and electron density and thus act as corrosion inhibitor and are the active centers for the process of adsorption on the metal surface. The inhibition efficiency conforms to the sequence  $O < N < S < P$  [2]. The use of corrosion inhibitors that are biodegradable, in addition to being ecologically acceptable, have made plant products good alternatives to synthesized organic compounds [3]. A number of investigations have been conducted on extracts of different parts of plant materials as potential mild steel corrosion inhibitors in acidic media, thus: cashew waste [4]; *Jatropha tanjorensis* stem bark [5], *Pentaclethra macrophylla* Benth roots [6]; *Cuminum cyminum* extract [7]; Papaya Leaves [8]; Cassava leaf [9]; water melon wastes [10]; coconut shell [11]; *Gmelina arborea* root extract [12] and *Vernonia amygdalina* extract [13].

Cola belongs to the family of *Sterculiaceae* [14]. The five species of edible cola nuts genus include: *C. nitida*, *C. acuminata*, *C. ballayi*, *C. verticillata* and *C. sphaerocarpa*. The latter three species are not known to be cultivated [15]. *C. nitida* is referenced to have phytochemicals that possess extensive conjugation with substantial  $\pi$  – electron systems and as well contain heteroatoms such as N, S, O or heterocyclic ring in their structures that serve as rally points for interaction with the metal surface [16; 1; 15; 17]. These characteristics enhance the potentials of the phytochemicals in the seeds extract of *C. nitida* to inhibit corrosion. Therefore, this research is inclined to investigating the corrosion inhibitory potentials of eco-friendly and less toxic isolated phytochemical from the seeds of *C. nitida*, as a green corrosion inhibitor.

## 2 EXPERIMENTAL DETAILS

### 2.1 PLANT MATERIAL AND EXTRACTION

Pods having fresh and mature seeds of *C. nitida* were harvested from a plantation in Mbikpong, a rural settlement in Ibesikpo/Asutan Local Government Area of Akwa Ibom State and identified by a taxonomist in the Department of Pharmacognosy and Natural Medicine, Faculty of Pharmacy, University

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of Uyo, Nigeria with a Voucher No. of UUPH 46. The pods were dissected with a sharp object and the seeds were collected and rinsed with water, cut into pieces, crushed and dried at room temperature ( $30 \pm 2$  °C) for 24 hours. It was then put through comparative extraction in which ethanol, acetone, methanol, ethyl acetate and dichloromethane (DCM) were used to macerate 400 g per solvent (2.5 L) in glass extraction jars for 72 hours and re-macerated for another 72 hours for maximum yield. Concentration of the filtrates was achieved using Grant GLS400 water-bath and a preliminary weight loss test was performed on the extracts and DCM extract was affirmed the most active fraction. This was weighed (85 g) and preserved in a refrigerator.

## 2.2 PURIFICATION OF DCM EXTRACT

Inherent components in DCM fraction of the extract was subjected to vacuum liquid chromatography (VLC) in order to enhance separation using the following solvent systems: 100 % hexane, 70% hexane: 30% DCM, 50% hexane: 50% DCM (re-macerated), 30% hexane: 70% DCM, 100% DCM, 70% DCM: 30% ethyl acetate, 50% DCM: 50% ethyl acetate, 30% DCM: 70% ethyl acetate and 100% ethyl acetate. Based on the spectra of thin layer chromatographic characteristics, the fractions obtained were pooled, thus:  $V_1 = 100\% \text{ hexane} + 70\% \text{ hexane:}30\% \text{ DCM}$  (38.7 g);  $V_2 = 50\% \text{ hexane:}50\% \text{ DCM}$  (61.5 g);  $V_3 = 50\% \text{ hexane:}50\% \text{ DCM} + 30\% \text{ hexane:}70\% \text{ DCM}$  (85.1 g);  $V_4 = 100\% \text{ DCM} + 70\% \text{ DCM:} 30\% \text{ ethyl acetate}$  (50.9 g); and  $V_5 = 50\% \text{ DCM:} 50\% \text{ ethyl acetate} + 30\% \text{ DCM:} 70\% \text{ ethyl acetate} + 100\% \text{ ethyl acetate}$  (87.3 g). Initial gravimetric tests on the pooled VLC fractions showed that corrosion inhibition efficiencies were in the following order:  $V_5$  (89%) >  $V_3$  (85 %) >  $V_1$  (83 %) >  $V_4$  (77 %) >  $V_2$  (71%).

Gradient elution with n-hexane containing increasing quantity of dichloromethane followed by increasing quantities of ethyl acetate and methanol (MeOH) was used on  $V_5$  (87.3 g) loaded on to silica gel column (C1) chromatography (Merck, 60-120 mesh). Eluates of 20 mL were obtained and examined on silica TLC plates (Merck, Germany) in DCM (100%) using p-anisaldehyde-sulphuric acid as spray reagent. Eighty-four (84) eluates were obtained and bulked into eight fractions [A =  $O_{1-12}$ (4.9 g), B=  $O_{13-23}$ (9.2 g), C =  $O_{24-31}$ (8.2 g), D =  $O_{32-59}$ (5.8 g), E =  $O_{60-69}$ (8.6 g), F =  $O_{70-77}$ (16.1 g), G =  $O_{78-81}$  (15.9 g) and H =  $O_{82-84}$ (10.2 g)] based on their TLC characteristics; silica gel solvent systems were: 100 % DCM, DCM (2): MeOH (8), DCM (8): MeOH (2) and Ethyl acetate(100%). Based on the TLC

characteristics, bulk fraction code named G (114.3 g) was subjected to further purification process by means of normal phase column (C2) chromatography and gradient-eluted with ethyl acetate containing increasing quantities of methanol. 20 mL each of the eluates were collected, monitored on silica TLC plates (Merck, Germany) in DCM (100%) and DCM (8): methanol (2) using p-anisaldehyde-sulphuric acid as spray reagent. Thirty-eight eluates were obtained and bulk into five sub-fractions, thus: A1 = O<sub>2-4</sub>(2.5 g), B1 = O<sub>5-7</sub> (1.3 g), C1 = O<sub>8-15</sub>(6.8 g), D1 = O<sub>16-35</sub>(4.1 g) and E1 = O<sub>36-38</sub>(2.1 g) based on their TLC characteristics. Sub-fraction D1 was colorless from column two (C2) gave single spot on TLC plate with R<sub>F</sub> of 0.82, confirmed pure and code named D1.

### 2.2.1 CHARACTERISATION OF ISOLATED PHYTOCHEMICAL

The isolated compound code named D1 (with single spot on TLC plate) was subjected to Fourier transform infra-red (FT-IR) spectroscopy, proton (<sup>1</sup>H), carbon-13 (<sup>13</sup>C) and distortionless enhancement by polarization transfer-135 (DEPT-135) nuclear magnetic resonance (NMR) spectroscopic studies.

### 2.3 MILD STEEL COUPONS

The mild steel sheet used was procured from John & Sons Nigeria Limited, Uyo, Akwa Ibom State. The steel which had the chemical composition (w/w%): C (0.12), Mn (0.85), S (0.06), P (0.05), Si (0.09) and Fe (98.83). Mechanically, the steel was pressed - cut into coupons of 3.0 cm x 2.0 cm x 0.05 cm. At the center of each coupon, a small hole was drilled to facilitate total immersion in the corrodent using glass rod and hook. These coupons were abraded with emery papers of different grades (P60X, 220 and 450) to obtain smooth finished surfaces. The coupons were rinsed with deionized water to get rid of organic impurities, washed in ethanol to remove grease and thereafter in acetone to remove oil. The cleaned coupons were preserved over calcium chloride in desiccators prior to use.

The dimension of the coupons for electrochemical experiment was 1.0 cm x 1.0 cm. They were also polished to a smooth finish.

### 2.4 WEIGHT LOSS MEASUREMENTS

Weight loss measurements were performed at total immersion condition of the coupons using glass rod and hook in 100 ml of 1M HCl solution containing D1 and blank, respectively, in a beaker maintained at 303 K. One coupon was placed in each beaker. Mild steel coupons were retrieved from the inhibited and blank solutions at intervals of two hours, immersed in 20% NaOH solution containing 200 gL<sup>-1</sup> of zinc dust, scrubbed with bristle brush, washed in distilled water, dried in acetone and reweighed [18] to determine weight loss with respect to time. This was done for five consecutive times giving a total of ten hours per temperature. The weight loss was taken to be the difference between the initial weight and the weight of the coupons at a given time for the chosen temperature range. The procedure was repeated at 313 K, 323 K and 333 K. From the data, corrosion rate (CR), surface coverage ( $\theta$ ) and percentage inhibition efficiency (IE %) were determined using the following equations [19].

$$CR = \frac{W_1 - W_2}{St} \quad (1)$$

where  $W_1$  is the mass of mild steel coupon before immersion,  $W_2$  is the mass of mild steel coupon after immersion, S is the total surface area of the coupon and t is the corrosion time in hours.

$$\theta = \frac{CR_0 - CR}{CR_0} \quad (2)$$

$$IE \% = \theta \times 100 \quad (3)$$

where  $CR_0$  and  $CR$  are corrosion rates of the mild steel coupon in 1.0 M HCl in the absence and presence of the extracts, respectively.

## 2.4 ELECTROCHEMICAL METHOD

Electrochemical experiments were conducted using the conventional three-electrode cell with a platinum counter electrode (CE) and a saturated calomel electrode (SCE) coupled to a fine lugging capillary as the reference electrode (RE). Ohmic contribution was minimized by placing the lugging capillary closed to the working electrode (WE) which the prepared mild steel coupon was presented in a square form of dimension 1.0 cm x 1.0 cm. The electrode was immersed in the test solution at an open circuit potential (OCP) for thirty minutes to stabilize before measurements could be taken. All electrochemical measurements were taken with the sample immersed in the blank, maximum and minimum concentrations of the test solutions at 303 K using Gamry 600+ Potentiostat/Galvanostat/ZRA Advanced Electrochemical System linked to a computer. The electrochemical set-up was sealed with

butyl stoppers where the electrodes and nitrogen pipe providing N<sub>2</sub> gas for deaeration could pass through. Prior to the electrochemical test, the glass cells were sterilized in an autoclave. Both the working electrode and the counter electrodes were further sterilized using 75% ethanol and kept under ultraviolet rays prior to the experiment.

#### 2.4.1 POTENTIODYNAMIC POLARIZATION MEASUREMENTS

The polarization curve was recorded under potentiodynamic polarization at the condition of air atmosphere and controlled by a personal computer. After the open circuit potential had been established, potentiodynamic polarization curves were obtained at a scan rate of 1mV<sup>-1</sup> in the potential range from -0.25 V to 0.25 V. Application of a small potential drop ( $\Delta E$ ) generates a corresponding current flow ( $\Delta I$ ) [20].

The corrosion rate of the system was determined using corrosion current density,  $i_{corr}$ . The linear portions of the anodic and cathodic Tafel curves were extrapolated to the corrosion potential ( $E_{corr}$ ) and the corrosion current ( $i_{corr}$ ) axes. All potentials were measured against SCE. The surface coverage ( $\theta$ ) and the percentage inhibition efficiency (% IE), were evaluated by the application of Equations (3.4) and (3.5) (Roy and Sukul, 2015), thus:

$$\theta = \frac{i_{corr}^o - i_{corr}}{i_{corr}^o} \quad 4$$

$$IE\% = \theta \times 100 \quad 5$$

**Comment [h7]:** Can we say that this equation is repeated and the same as equation number 3 & 7?

The values of corrosion current density in the absence and presence of inhibitor are given by  $i_{corr}^o$  and  $i_{corr}$ , respectively [21; 22].

#### 2.4.2 ELECTROCHEMICAL IMPEDANCE SPECTROSCOPIC MEASUREMENTS

The measurements of electrochemical impedance spectroscopy were carried out over a frequency range of 10 Hz to 100,000 Hz at 303 K using amplitude of 5 mV RMS peak-to-peak with an alternating current signal at the open circuit potential and at air atmosphere. The impedance data were recorded using Nyquist plots and the charge transfer resistance,  $R_{ct}$ , was obtained from the diameter of

the semicircle in the plot. The percentage inhibition efficiency (% IE) was calculated from  $R_{ct}$  values obtained from the impedance measurements according to the models bellows, thus:

$$\theta = \frac{R_{ct} - R_{ct}^0}{R_{ct}} \quad 6$$

$$IE\% = \theta \times 100 \quad 7$$

charge transfer resistances in the presence and absence of the inhibitor are given by  $R_{ct}$  and  $R_{ct}^0$ , respectively [23; 24].

## 2.5 SCANNING ELECTRON MICROSCOPY (SEM)

Samples were ground with SiC paper having 1000 grit, washed in distilled water and then in ethanol, dried in compressed air and the placed in a fume hood. Morphological survey was conducted on 1 cm x 1 cm area of the surface of mild steel coupons before and after immersion in both blank and test solutions. Thermo Fisher Scientific Prisma E scanning electron microscopy was used. Measurements were taken at various magnifications and the appropriate ones selected.

## 2.6 QUANTUM CHEMICAL STUDIES

Corrosion inhibition mechanism of D1 on mild steel in hydrochloric solution was studied using Gaussian 09 and density functional theory (DFT) with the basis set of B3LYP/6 – 31G (d,p) in aqueous phase. Possible physical properties of D1 that make it prone to corrosion inhibition were evaluated using theoretical study.

## 3 RESULTS AND DISCUSSION

### 3.1 CHARACTERISATION OF D1

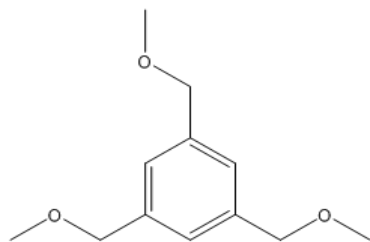
D1 was obtained in a substantial quantity (4.1 g) as a colorless crystalline solid soluble in dimethyl sulphur (II) oxide (DMSO) solvent. Spectroscopic results on D1 revealed, thus: FT- IR (KBr,  $\text{cm}^{-1}$ ): 3153, 3116, 3015 (C-H aromatic), 2829 (C-H aliphatic), 1655 (C=C), 1454 ( $\text{CH}_3$  deformation), 1223 (C-O stretching).

$^1\text{H}$  NMR (60 MHz,  $\text{DMSO}-d_6$ ):  $\delta$  7.95 (s, 3H, Ar-H), 3.85 (s, 6H, 3 x  $\text{CH}_2\text{OCH}_3$ ), 1.24 (s, 9H, 3 x  $\text{CH}_2\text{OCH}_3$ ) ppm.

C-13 NMR (15 MHz,  $\text{DMSO}-d_6$ ): No information was deduced from the spectra due to absence of visible compound peaks.

DEPT-135 (15 MHz, DMSO- $d_6$ ): No information was deduced from the spectra due to absence of visible compound peaks.

Based on the available pieces of information, the following structure was elucidated, thus:



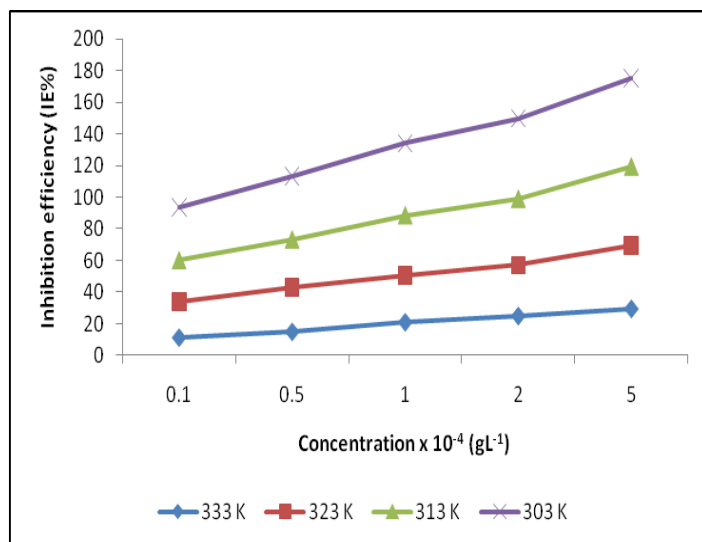
**Figure 1:** 1, 3, 5-Trimethoxymethylbenzene (TMMB)

Therefore, D1 was elucidated and identified as 1, 3, 5-trimethoxymethylbenzene abbreviated TMMB.

## 3.2 WEIGHT LOSS RESULTS

### 3.2.1 EFFECT OF 1, 3, 5-TRIMETHOXYMETHYLBENZENE (TMMB) CONCENTRATION ON INHIBITION EFFICIENCY

Variation of inhibition efficiency with respect to the concentration of TMMB in 1 M HCl solution is presented in Figure 2. It is observed that the inhibition efficiency increased with increase in TMMB concentration. The highest inhibition efficiency obtained at 303 K was 55.83% at  $5.0 \times 10^{-4} \text{ g L}^{-1}$  (Table 1). This implies that there is an interaction between TMMB molecules and the steel surface. The data indicate that there is a higher degree of surface coverage of mild steel surface with increase in the concentration of TMMB, thereby impeding the actions of aggressive ions and of course inhibit corrosion [25].



**Figure 2:** Effect of 1, 3, 5-trimethoxymethylbenzene (TMMB) on the corrosion inhibition efficiency of mild steel at different temperatures.

**Table 1:** Calculated values of weight loss, corrosion rate and inhibition efficiency for mild steel corrosion in 1 M HCl in the absence and presence of different concentrations of TMMB

Inhibitor Conc. ( $\times 10^{-4}$ )	Weight loss (g)				Corrosion rate ( $\text{mg cm}^{-2} \text{hr}^{-1}$ )				Inhibition efficiency (%)			
	303 K	313 K	323 K	333 K	303 K	313 K	323 K	333 K	303 K	313 K	323 K	333 K
Blank	0.150	0.248	0.317	0.421	2.40	3.97	5.07	6.74	-	-	-	-
0.1 $\text{g L}^{-1}$	0.100	0.184	0.214	0.361	1.60	2.94	3.92	5.94	33.33	25.94	22.68	11.42
0.5 $\text{g L}^{-1}$	0.090	0.174	0.198	0.358	1.44	2.78	3.64	5.73	40.00	29.97	28.21	14.99
1.0 $\text{g L}^{-1}$	0.082	0.154	0.184	0.333	1.31	2.46	3.57	5.33	45.42	38.04	29.59	20.92
2.0 $\text{g L}^{-1}$	0.074	0.144	0.156	0.317	1.18	2.30	3.44	5.01	50.83	42.07	32.15	24.78
5.0 $\text{g L}^{-1}$	0.066	0.124	0.132	0.238	1.06	1.98	3.04	4.77	55.83	50.13	50.13	29.23

### 3.2.2 EFFECT OF TEMPERATURE ON INHIBITION EFFICIENCY

Table 1 displays a pattern of corrosion rate (CR) considering total immersion time of 10 hours in the absence and presence of TMMB at different temperatures. It is seen that CR increases with increase in temperature and decreases with increase in the concentration as it varies, thus:  $1.06 \text{ mg cm}^{-2} \text{hr}^{-1}$  at 303K,  $1.98 \text{ mg cm}^{-2} \text{hr}^{-1}$  at 313 K,  $3.04 \text{ mg cm}^{-2} \text{hr}^{-1}$  at 323 K and  $4.77 \text{ mg cm}^{-2} \text{hr}^{-1}$  at 333 K considering  $5.0 \times 10^{-4} \text{ g L}^{-1}$  of TMMB (Table 1). This is because TMMB molecules readily acquire more energy at elevated temperatures and overcame the activation energy barrier than at lower temperatures (Abakedi and Asuquo, 2016). Besides, weight loss trend is seen to decrease as the concentration of TMMB increases. This is linked to the protection of mild steel as the degree of surface coverage ( $\theta$ ) increased proportion to the concentration of TMMB. At the concentration of  $0.1 \times 10^{-4} \text{ g L}^{-1}$  and  $5.0 \times 10^{-4} \text{ g L}^{-1}$  at

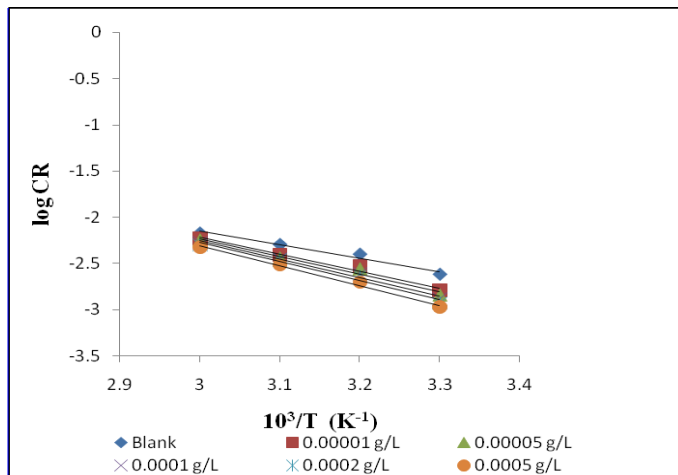
303 K, the weight loss by TMMB were 0.100 g and 0.066 g, respectively (Table1). [26] had reported a similar trend.

Corrosion rate, CR, evaluated from the weight loss measurements in acidic medium is related to temperature by the Arrhenius equation:

$$\log CR = \log A - \frac{E_a}{2.303RT} \quad (8)$$

where  $E_a$  is the energy of activation, A is the pre-exponential constant of Arrhenius, R is the universal gas constant and T is the absolute temperature. Energies of activation were evaluated from the slopes of the plot of  $\log CR$  versus  $\frac{1}{T}$  (Figure 3) and presented in Table 2. It is obvious that the  $E_a$  values in the presence of TMMB are higher than the  $E_a$  of the blank. These values reveal that the presence of TMMB raised the energy of activation of the metal dissolution reaction. The higher values of energy of activation in the presence of TMMB compared to the blank and the decrease in the percentage inhibition efficiency with increase in temperature are pointers to physical adsorption process [27]. Therefore, the mechanism of adsorption of TMMB on mild steel surface is by physical adsorption.

Comment [h8]: ITS Arrhenius like equation



**Comment [h9]:** Re draw the chart make the y-axis between -2 \_ -3 and define each axis

**Figure 3:** Arrhenius plot for mild steel coupons in 1 M HCl solution in the absence and presence of different concentrations of TMMB.

**Table 2:** Thermodynamic parameters for mild steel corrosion in 1 M HCl solution in the absence and presence of different concentrations of TMMB at 303 – 333 K.

Inhibitor	Concn.	$E_a$	$\Delta H_{ads}^{\circ}$	$-\Delta S_{ads}^{\circ}$
	M x 10 <sup>-4</sup>	(kJ mol <sup>-1</sup> )	(kJ mol <sup>-1</sup> )	(J K <sup>-1</sup> mol <sup>-1</sup> )
Blank		27.95	25.16	211.63
TMMB	0.1	35.04	32.45	190.94
	0.5	36.76	34.06	186.44
	1.0	38.10	35.46	182.68
	2.0	39.63	36.80	179.03
	5.0	40.78	38.47	174.66

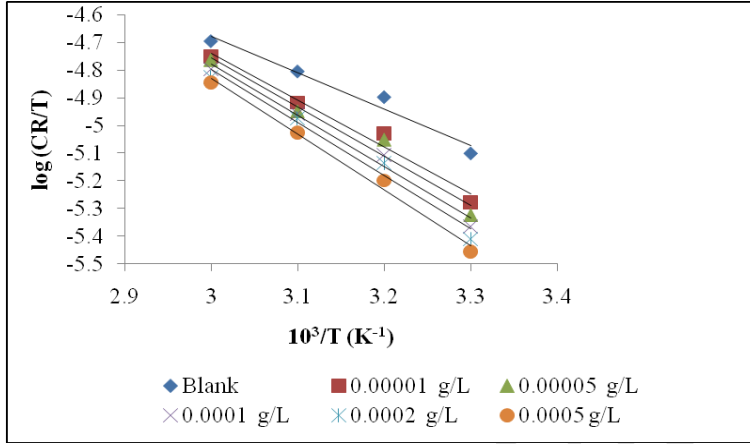
The values of enthalpy change,  $\Delta H_{ads}^{\circ}$ , and entropy change of adsorptions,  $\Delta S_{ads}^{\circ}$ , were determined from the transition state equation [28] as stated in equation (9), thus:

$$CR = \frac{RT}{Nh} \exp\left(\frac{\Delta S_{ads}^{\circ}}{R}\right) \exp\left(-\frac{\Delta H_{ads}^{\circ}}{RT}\right) \quad (9)$$

where N is the Avogadro's number, h is the Planck's constant, R is the universal gas constant and T is the absolute temperature. Equation (9) is rearranged as follows:

$$\ln\left(\frac{CR}{T}\right) = \left[\ln\left(\frac{R}{Nh}\right) + \frac{\Delta S_{ads}^{\circ}}{R}\right] - \frac{\Delta H_{ads}^{\circ}}{RT} \quad (10)$$

Linear plot of  $\log\left(\frac{CR}{T}\right)$  versus  $\frac{1}{T}$  (Figure 4) with slope and intercept of  $-\frac{\Delta H_{ads}^{\circ}}{R}$  and  $\left[\ln\left(\frac{R}{Nh}\right) + \frac{\Delta S_{ads}^{\circ}}{R}\right]$ , respectively, were obtained. The slopes were used to determine the enthalpy change of adsorption ( $\Delta H_{ads}^{\circ}$ ) and intercepts were used to obtain the entropy change of adsorption ( $\Delta S_{ads}^{\circ}$ ) as presented in Table 3. The positive values of  $\Delta H_{ads}^{\circ}$  connote that the adsorption process of TMMB on to the surface of mild steel accompanied heat adsorption and it is endothermic in nature. The negative values of  $\Delta S_{ads}^{\circ}$  imply that the adsorption process is linked to a decrease in entropy and this is paramount to the adsorption of TMMB on to the surface of mild steel surface [29; 30]. Similar result had been reported by [25].



**Figure 4:** Transition State plot for mild steel coupons in 1M HCl in the absence and presence of different concentrations of TMMB.

### 3.2.3 ADSORPTION CONSIDERATION

It is imperative to evaluate an experimental best fit adsorption isotherm to the surface coverage data in order to assess the interactions between the molecules of TMMB and the mild steel surface. The best fit was obtained with Temkin adsorption isotherm given as equation (6) [31; 32; 33].

$$e^{f\theta} = K_{ads} C \quad 11$$

where  $\theta$  is the surface coverage,  $K_{ads}$  is the equilibrium constant of adsorption process,  $C$  is the inhibitor concentration and  $f$  is the factor of energetic inhomogeneity describing the molecular interactions in the adsorption layer ( $f = -2a$ ) [34].

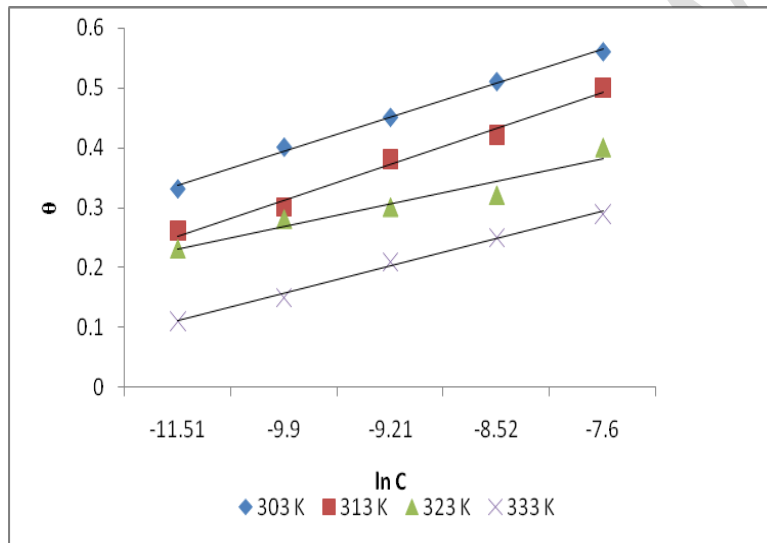
$K_{ads}$  is related to standard free energy of adsorption,  $\Delta G_{ads}$ , through the equation (8) [35].

$$\log K_{ads} = 1.744 - \frac{\Delta G_{ads}^{\circ}}{2.303 RT} \quad 12$$

A linear plot of surface coverage ( $\theta$ ) versus logarithm of inhibitor concentration ( $\log C$ ) with correlation coefficients ( $R^2$ ) within the range  $0.930 \leq R^2 \leq 0.996$  obtained for TMMB (Figure 5) indicates that TMMB adsorption best fits to Temkin isotherm. The values of  $K_{ads}$  signify the strength of adsorption between adsorbate and adsorbent. Large values of  $K_{ads}$  reflect efficient adsorption and hence better inhibition efficiency [36]. The values of  $K_{ads}$  from Table 3 are very low and this implies weak interaction (physical

adsorption) between the molecules of TMMB and the mild steel surface. The negative values of 'a' connote existence of attraction in the adsorption layer [37].

The negative values of Gibbs' free energy change of adsorption,  $\Delta G_{ads}$ , (Table 3) imply that the adsorption process of TMMB on to the surface of mild steel was spontaneous.



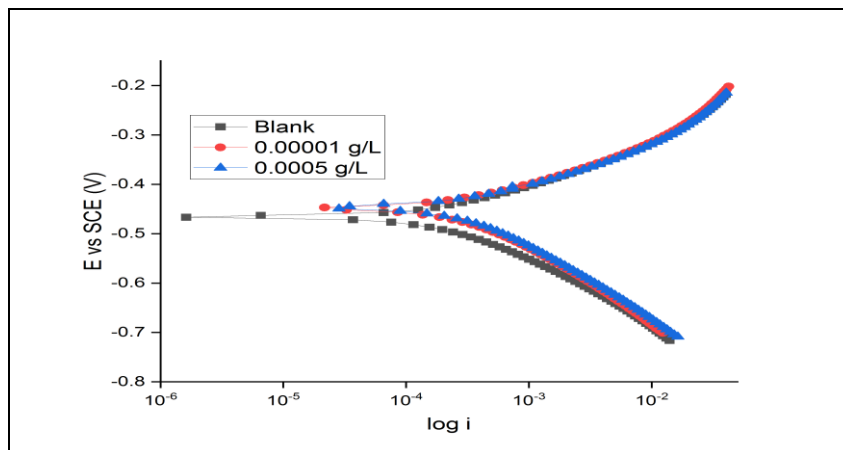
**Figure 5:** Temkin adsorption isotherm for mild steel corrosion in 1 M HCl solution in the absence and presence of different concentrations of TMMB.

**Table 3:** Some adsorption parameters from Temkin isotherm for mild steel coupons in 1 M HCl containing different concentrations of TMMB at 303 – 333 K.

Inhibitor	Temp. (K)	Intercept	Slope	K (mol/L)	f	- a	-ΔG (kJ/mol)	R <sup>2</sup>
TMMB	303	0.279	0.057	133.44	17.54	8.77	22.45	0.996
	313	0.192	0.060	24.55	16.67	8.33	18.78	0.986
	323	0.192	0.038	156.56	26.32	13.16	24.36	0.930
	333	0.064	0.046	4.02	21.74	10.87	14.97	0.994

### 3.3 POTENTIODYNAMIC POLARIZATION MEASUREMENTS

The presence of TMMB in the test solutions pushed the cathodic and anodic locations to the minimum current magnitudes. This means that corrosion at these locations was retarded by TMMB. Polarization parameters such as corrosion current densities ( $i_{corr}$ ), corrosion potential ( $E_{corr}$ ), cathodic Tafel slope ( $\beta_c$ ) and anodic Tafel slope ( $\beta_a$ ) were generated from the Tafel plot (Figure 6) and used in the calculation of percentage inhibition efficiency (IE%) as presented in Table 4 at different concentrations of TMMB. The values of  $\beta_c$  and  $\beta_a$  were largely affected by the inclusion of TMMB in the blank. When the shift in  $E_{corr}$  is higher than 85 mV, the inhibitor can be grouped as either anodic or cathodic, but if the shift is lower than 85 mV, the inhibitor is categorised mixed-type [38]. The  $E_{corr}$  values as presented in Table 4 are far less than the threshold value of 85 mV. This confirms that TMMB is a mixed-type corrosion inhibitor with the inhibition of both cathodic and anodic reactions. Similar results had been reported by [39]. Percentage inhibition efficiencies of 67% and 48% were recorded at  $5.0 \times 10^{-4}$  g L<sup>-1</sup> and  $0.1 \times 10^{-4}$  g L<sup>-1</sup> concentrations of TMMB, respectively. This means that  $i_{corr}$  reduces considerably with respect to increase in the concentration of TMMB in the blank (Table 4). Therefore, percentage inhibition efficiency improves upon increase in the concentration of the inhibitor [40].



**Figure 6:** Polarization curve of mild steel in 1.0 HCl solution in the absence and presence of different concentrations of TMMB at 303 K.

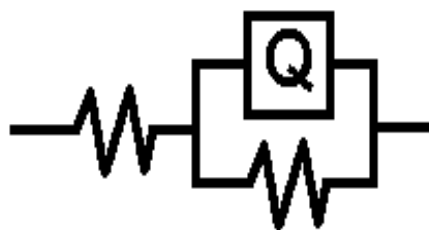
**Table 4:** Potentiodynamic polarization parameters from Tafel plots for mild steel immersed in 1.0 M HCl solution in the absence and presence of different concentrations of TMMB at 303 K.

Inhibitor	Concn. (g/L)	$-E_{\text{corr}}$ (mV vs SCE)	$i_{\text{corr}}$ ( $\mu\text{A}/\text{cm}^2$ )	$\beta_c$ (mV dec $^{-1}$ )	$\beta_a$ (mV dec $^{-1}$ )	IE%
	Blank	453.11	365.01	154	65	-
TMMB	$0.1 \times 10^{-4}$	458.66	190.05	140	98	48
	$5.0 \times 10^{-4}$	461.54	118.65	124	86	67

### 3.4 ELECTROCHEMICAL IMPEDANCE SPECTROSCOPY

A one-time constant phase element (CPE) circuit (Figure 7) was used to analyze the electrochemical impedance data generated from a sample area of  $1.0 \text{ cm}^2$  at chi squared error of  $6.081 \times$

$10^{-6}$ , relative standard error ranging from 0.084 to 0.788 and the measurement errors in impedance data  $\leq 0.2466$ . This was achieved using the ZSimpwin software application.



**Figure 7:** One time constant phase element (CPE) circuit

Impedance measurements were performed to evaluate the characteristics and kinetics of the metal/solution interface and how the corrosion reactions are obstructed by the inhibitor [41].

A plot of  $Z'_{im}$  versus  $Z'_{Re}$  (Nyquist) of mild steel immersed in the absence and presence of TMMB is presented in Figure 8. It is observed that impedance of the mild steel increases with increase in the concentration of TMMB. The plot shows one incomplete capacitive loop corresponding to the charge transfer reaction which is dependent on either electron transfer at the metal surface or on the electron conduction through the surface film [42]. It is observed that the change in concentration of TMMB had no effect on the shape of the impedance behavior and this suggests that TMMB exhibits similar corrosion inhibition mechanism on mild steel at various concentrations.

The impedance plots (Figure 8) are not perfect semicircles due to frequency dispersion caused by coarseness and surface fractions of the metal [43]. The impedance of the CPE is obtained according to equation (10) [44; 45], thus:

$$Z_{CPE} = Y_0^{-1} (j\omega)^{-n} \quad 13$$

where  $Y_0$  ( $1.182 \times 10^{-3}$  at  $5.0 \times 10^{-4} \text{ gL}^{-1}$  and  $1.250 \times 10^{-3}$  at  $0.1 \times 10^{-4} \text{ gL}^{-1}$  of TMMB) and  $n$  ( $0.817 \leq n \leq 0.830$ ) are associated with CPE (Table 6).  $j^2 = -1$  regarded as the imaginary axis,  $\omega$  stands for angular

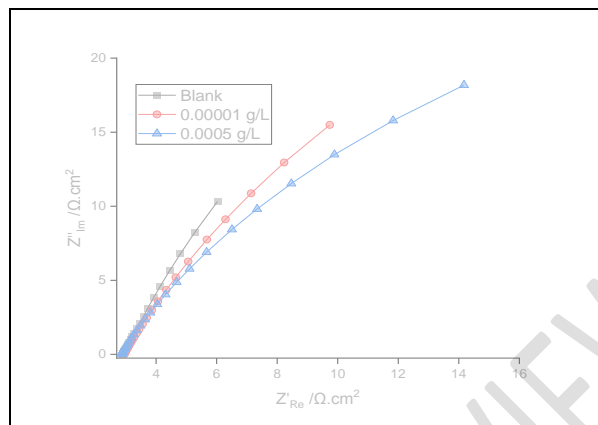
frequency in  $\text{rads}^{-1}$ ,  $W$  denotes the impedance factor and  $n$  is the considered shifting factor which explains the surface heterogeneity, impurities and surface fractions of the metal [39].

Chloride ions from the electrolyte (acid used) adsorb onto the metal surface and the presence of TMMB in test solution gradually displaces the chloride ions and their presence at relatively low inhibitor concentration may induce adsorption. More so, chloride ions play a role in the adsorption of protonated inhibitor molecules by specially adsorbing to the metal surface, thus creating a negative charge towards the solution [46]

Introduction of TMMB in the acid solution promoted the radius of the imperfect loop which reflects corrosion inhibition and corresponds to the charge transfer resistance,  $R_{ct}$ , of the mild steel surface. Diameter of the depressed incomplete semicircles at high frequency is often regarded as the charge transfer resistance,  $R_{ct}$ . The percentage inhibition efficiency,  $IE\%$ , of TMMB for the mild steel electrode is evaluated using equation (11), thus:

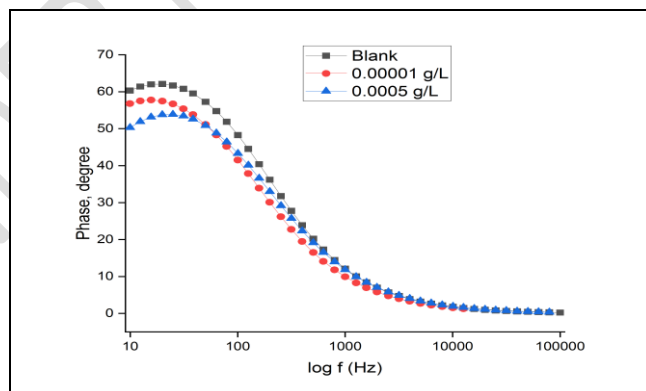
$$IE \% = \frac{R_{ct} - R_{ct}^0}{R_{ct}} \times 100 \quad 14$$

where charge transfer resistances in the presence and absence of TMMB are given by  $R_{ct}$  and  $R_{ct}^0$ , respectively [24]. The value of  $R_{ct}$  increases in direct proportion to the concentration of TMMB, with percentage inhibition efficiency of 58% at  $5.0 \times 10^{-4} \text{ gL}^{-1}$  at 303 K (Table 6).



**Figure 8:** Nyquist plot of mild steel in 1.0 M HCl solution in the absence and presence of different concentrations of TMMB at 303 K.

Bode plot of phase angle against logarithm of frequency ( $\log f$ ) in the absence and presence of TMMB is presented in Figure 9 and phase angle peaks at different frequencies of the corrosion media at 303 K are obvious. Asymmetric overlap of peaks for each concentration of TMMB are observed at 37.99 Hz as  $52.4^\circ$  for  $5.0 \times 10^{-4} \text{ gL}^{-1}$ , at 29.12 Hz as  $57.1^\circ$  for  $0.1 \times 10^{-4} \text{ gL}^{-1}$  and at 31.3 Hz as  $62.3^\circ$  for the blank (Figure 8), respectively. This implies that TMMB inhibits mild steel corrosion in acidic media.



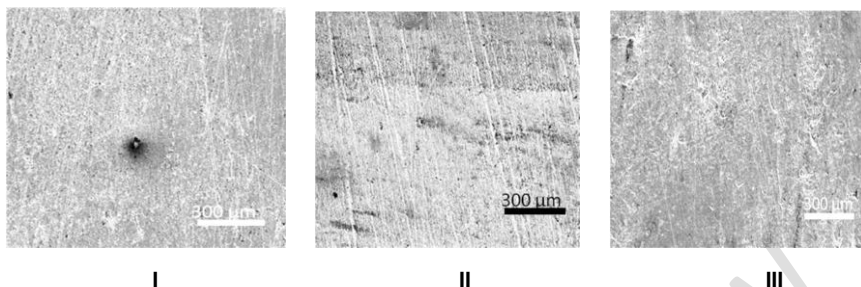
**Figure 9:** Bode Plot of mild steel in 1.0 M HCl solution in the absence and presence of different concentrations of TMMB at 303 K.

**Table 5:** Some Electrochemical Impedance Parameters of Mild Steel in 1M HCl Solutions in the absence and presence of TMMB at 303 K.

TMMB Concn.	$R_s$	$R_{ct}$	Q		IE %
	( $\Omega \cdot \text{cm}^{-2}$ )	( $\Omega \cdot \text{cm}^{-2}$ )	$Y_0 \times 10^{-3}$	n	
Blank	2.914	76.550	0.831	0.856	-
$0.1 \times 10^{-4} \text{ g L}^{-1}$	2.947	125.490	1.250	0.817	40
$5.0 \times 10^{-4} \text{ g L}^{-1}$	2.851	182.260	1.182	0.830	58

### 3.5 SCANNING ELECTRON MICROSCOPY (SEM)

Micrographs of the mild steel coupons placed in 1 M hydrochloric acid solution in the absence and presence of 1, 3, 5-Trimethoxymethylbenzene (TMMB) were taken at a magnification of 300  $\mu\text{m}$  and are presented in Figure 10 (I – III). Figure 9a exposes localized pitting of the coupon in blank. The aggressive nature of  $\text{Cl}^-$  and  $\text{H}^+$  ions in the interfacial dissolution layer of the steel is responsible for the rough morphology of mild steel in the blank [41]. Similar trend of results were obtained by other workers [47; 25]. The mild steel coupons in the presence of TMMB at  $0.1 \times 10^{-4} \text{ gL}^{-1}$  (Figure 10 II) and at  $5.0 \times 10^{-4} \text{ gL}^{-1}$  (Figure 10 III) are smoother than that of the blank. The differences in textural properties of the micrographs arise from the formation of dense film at the metal-solution interface by TMMB, hence a reduction in the substrate (mild steel) deterioration rate. Consequently, TMMB at  $5.0 \times 10^{-4} \text{ gL}^{-1}$  provided a better protection to the mild steel than TMMB at  $0.1 \times 10^{-4} \text{ gL}^{-1}$  as seen in the micrographs (Figure 10: I – III).



**Figure 10:** SEM image of mild steel surface exposed to (I) 1.0 M HCl solution (blank), (II) 1.0 M HCl containing  $0.1 \times 10^{-4} \text{ gL}^{-1}$  of TMMB, and (III) 1.0 M HCl containing  $5.0 \times 10^{-4} \text{ gL}^{-1}$  of TMMB.

### 3.6 QUANTUM CHEMICAL STUDIES

Optimized structure of 1, 3, 5-trimethoxymethylbenzene (TMMB), lowest unoccupied molecular orbital,  $E_{\text{LUMO}}$  and highest occupied molecular orbital,  $E_{\text{HOMO}}$  are presented in Figure 11 (I - III). Quantum chemical assessment reveals that frontier molecular orbital energies,  $E_{\text{HOMO}}$  and  $E_{\text{LUMO}}$ , are necessary parameters to determine the reactivity of chemical species by theoretical technique [42]. Quantum chemical parameters were calculated using equations (15 - 19) and are presented in Table 6. Information on the ability of a molecule to donate electrons to appropriate acceptor molecule with empty molecular orbital is obtained from  $E_{\text{HOMO}}$  and it is -6.570 eV with TMMB. The inhibition efficiency of an inhibitor is linked to the value of  $E_{\text{HOMO}}$  such that high value of  $E_{\text{HOMO}}$  reflects on high inhibition efficiency [48]. The energy gap,  $\Delta E_{\text{gap}}$ , determines the reactivity of molecules. Inhibitors with low  $\Delta E_{\text{gap}}$  should have good inhibition efficiencies because the energy required to remove the valence electron in a molecule is low and TMMB with  $\Delta E_{\text{gap}}$  (6.248 eV) is an average corrosion inhibitor. The number of electrons transferred ( $\Delta N$ ) reveals that TMMB ( $\Delta N = 0.220$ ). In view of the positive values of  $\Delta N$ , it implies that electrons will be donated to the metal surface by TMMB. It is important to note that in as much as the value of  $\Delta N$  is less than 3.6, the inhibition efficiency increases with the ability of the molecules to donate electrons to the metal surface [49]. The electronegativity,  $\chi$ , of TMMB is 3.446 eV is lower than the work function of Fe(110) surface given as 4.82 eV and this points to the fact that electrons migrate from the inhibitor to the

Fe surface. Since the value of  $\chi$  for TMMB is the low, movement of electrons from TMMB to the Fe surface is to be expected.

Global hardness,  $\eta$ , and global softness,  $\sigma$ , are used to define how an atom resists the deformation of its electron cloud and they are  $3.124 \text{ eV}$  and  $0.320 \text{ eV}^{-1}$ , respectively, implying that TMMB is a hard molecule. Therefore, TMMB is an average inhibitor from quantum studies.

$$\text{Energy gap, } \Delta E = E_{LUMO} - E_{HOMO} \quad 15$$

$$\text{Electronegativity, } \chi = -\frac{1}{2} (E_{HOMO} + E_{LUMO}) \quad 16$$

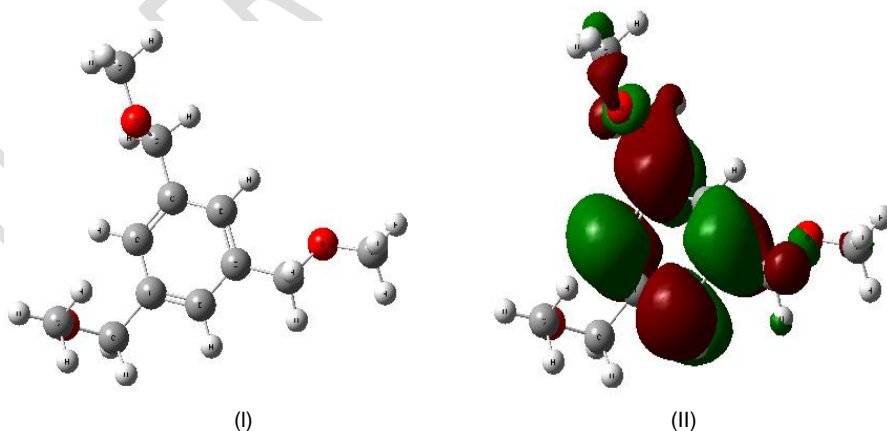
$$\text{Global hardness, } \eta = -\frac{1}{2} (E_{LUMO} - E_{HOMO}) \quad 17$$

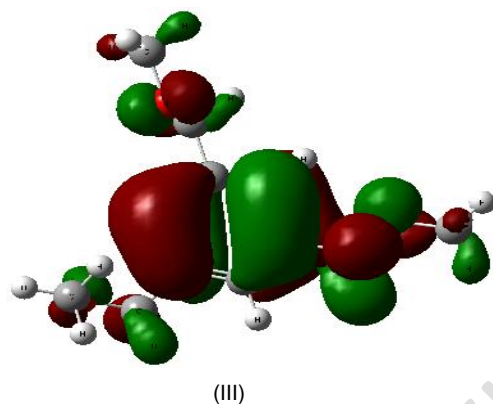
$$\text{Global softness, } \sigma = -\frac{1}{\eta} \quad 18$$

Number of transported electrons between metal surface and the inhibitor,

$$\Delta N = \frac{(\phi - \chi_i)}{2\eta_i} \quad 19$$

where  $\phi$  ( $4.82 \text{ eV}$ ) is the work function of iron [Fe(110)], a modeled surface configuration of iron used for the evaluation.





**Figure 11:** Structure of 1, 3, 5-Trimethoxymethylbenzene (TMMB): (I) optimized, (II) LUMO and (III) HOMO.

**Table 6:** Quantum chemical parameters for TMMB at geometry optimization

Inhibitor	$E_{\text{HOMO}}$ (eV)	$E_{\text{LUMO}}$ (eV)	$\Delta E_{\text{gap}}$ (eV)	IE (eV)	EA (eV)	$\chi$ (eV)	$\eta$ (eV)	$\sigma$ (eV <sup>-1</sup> )	$\Delta N$
TMMB	-6.570	-0.322	6.248	6.570	0.322	3.446	3.124	0.320	0.220

#### 4 CONCLUSION

The study reveals that 1, 3, 5 – trimethoxymethylbenzene (TMMB), a phytochemical isolated from fresh and mature seeds of *Cola nitida* inhibited the corrosion of mild steel in 1 M HCl solution. Inhibition efficiency of TMMB was found to increase with increase in concentration, but decreased with increase in temperature. Adsorption of TMMB onto the mild steel surface fitted to Temkin isotherm. The negative values of  $\Delta G_{\text{ads}}^{\circ}$  are indicative of the spontaneity of the corrosion inhibition process and the endothermic nature of TMMB adsorption process is confirmed by the positive values of  $\Delta H_{\text{ads}}^{\circ}$ . The negative values of  $\Delta S_{\text{ads}}^{\circ}$  indicate less disorder of the extract molecules on the mild steel surface and hence a stronger

adsorption of TMMB molecules on mild steel surface. TMMB is a mixed-type inhibitor based on its  $E_{corr}$  value. The inhibition efficiency obtained from weight loss, potentiodynamic polarization, electrochemical impedance spectroscopy and quantum chemical studies were in agreement. Dense film formed were seen in the micrographs of the steel and it was responsible for retarding the migration of aggressive ions in the corroder and thereby inhibiting mild steel corrosion in hydrochloric acid solution. Therefore, TMMB is concluded to be an average corrosion inhibitor of steel in hydrochloric acid solution.

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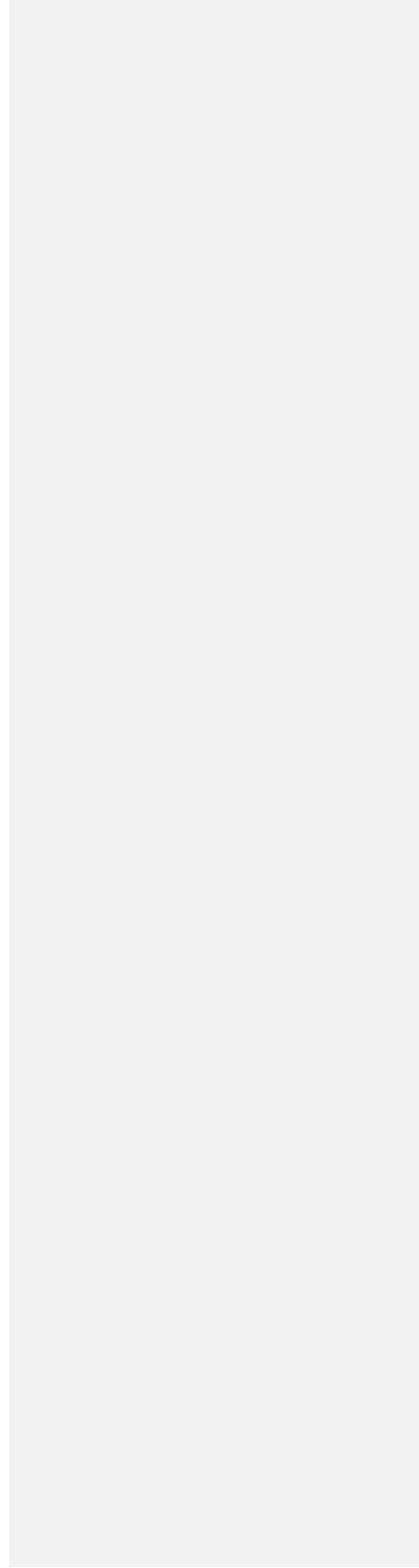
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